



Advanced Higher
course specification



Advanced Higher Chemistry, SCQF Level 7 Course Specification

Course code:	C813 77
Course assessment code:	X813 77
SCQF:	level 7 (32 SCQF credit points)
Valid from:	session 2026–27

This document provides detailed information about the course and course assessment to ensure consistent and transparent assessment year on year. It describes the structure of the course and the course assessment in terms of the skills, knowledge and understanding that are assessed.

This document is for teachers and lecturers and contains all the mandatory information required to deliver the course.

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Course overview

This course consists of 32 SCQF credit points, which includes time for preparation for course assessment. The notional length of time for candidates to complete the course is 160 hours.

The course assessment has two components.

Component	Marks	Scaled mark	Duration
Component 1: question paper	100	not applicable	2 hours and 30 minutes
Component 2: project	25	34	see 'Course assessment' section

Recommended entry	Progression
<p>Entry to this course is at the discretion of the centre.</p> <p>Candidates should have achieved the Higher Chemistry course or equivalent qualifications and/or experience prior to starting this course.</p>	<ul style="list-style-type: none">◆ a Higher National Diploma (HND), or degree in Chemistry or a related area, such as medicine, law, dentistry, veterinary medicine, engineering, environmental and health sciences◆ a career in a Chemistry-based discipline or related area such as renewable energy development, engineering, technology, pharmaceuticals, environmental monitoring, forensics, research and development, oil and gas exploration, management, civil service and education, or in a wide range of other areas◆ further study, employment and/or training

Conditions of award

The grade awarded is based on the total marks achieved across both course assessment components.

Course rationale

National Courses reflect Curriculum for Excellence values, purposes and principles. They offer flexibility, provide time for learning, focus on skills and applying learning, and provide scope for personalisation and choice.

Every course provides opportunities for candidates to develop breadth, challenge and application. The focus and balance of assessment is tailored to each subject area.

Chemistry, the study of matter and its interactions, plays an increasingly important role in most aspects of modern life. This course allows candidates to develop a deep understanding of the nature of matter, from its most fundamental level to the macroscopic interactions driving chemical change.

Candidates develop their abilities to think analytically, creatively, and independently to make reasoned evaluations, and to apply critical thinking in new and unfamiliar contexts to solve problems. The course offers candidates' flexibility and personalisation as they decide the choice of topic for their project.

Purpose and aims

The course builds on the knowledge and skills developed by candidates in the Higher Chemistry course and continues to develop their curiosity, interest and enthusiasm for chemistry in a range of contexts. Skills of scientific inquiry and investigation are developed throughout the course.

The course offers opportunities for collaborative and independent learning set within familiar and unfamiliar contexts, and seeks to illustrate and emphasise situations where the principles of chemistry are used and applied in everyday life.

Candidates develop important skills relating to chemistry, including developing scientific and analytical thinking skills and making reasoned evaluations.

The course aims to:

- ◆ develop a critical understanding of the role of chemistry in scientific issues and relevant applications, including the impact these could make in society and the environment
- ◆ extend and apply skills, knowledge and understanding of chemistry
- ◆ develop and apply the skills to carry out complex practical scientific activities, including the use of risk assessments, technology, equipment and materials
- ◆ develop and apply scientific inquiry and investigative skills, including planning and experimental design
- ◆ develop and apply analytical thinking skills, including critical evaluation of experimental procedures in a chemistry context
- ◆ extend and apply problem-solving skills in a chemistry context
- ◆ further develop an understanding of scientific literacy, using a wide range of resources, in order to communicate complex ideas and issues and to make scientifically informed choices
- ◆ extend and apply skills of autonomous working in chemistry

Who is this course for?

The course is suitable for candidates who are secure in their attainment of Higher Chemistry or equivalent qualifications. It is designed for candidates who can respond to a level of challenge, especially those considering further study or a career in chemistry and related disciplines.

The course emphasises practical and experiential learning opportunities, with a strong skills-based approach to learning. It takes account of the needs of all candidates, and provides sufficient flexibility to enable candidates to achieve in different ways.

Course content

The course content includes the following areas of chemistry:

Inorganic chemistry

The topics covered are:

- ◆ electromagnetic radiation and atomic spectra
- ◆ atomic orbitals, electronic configurations and the periodic table
- ◆ transition metals

Physical chemistry

The topics covered are:

- ◆ chemical equilibrium
- ◆ reaction feasibility
- ◆ kinetics

Organic chemistry and instrumental analysis

The topics covered are:

- ◆ molecular orbitals
- ◆ synthesis
- ◆ stereo chemistry
- ◆ experimental determination of structure
- ◆ pharmaceutical chemistry

Researching chemistry

The topics covered are:

- ◆ common chemical apparatus
- ◆ skills involved in experimental work
- ◆ stoichiometric calculations
- ◆ gravimetric analysis
- ◆ volumetric analysis
- ◆ practical skills and techniques

Skills, knowledge and understanding

Skills, knowledge and understanding for the course

The following provides a broad overview of the subject skills, knowledge and understanding developed in the course:

- ◆ extending and applying knowledge of chemistry to new situations, interpreting and analysing information to solve complex problems
- ◆ planning and designing chemical experiments/investigations, including risk assessments, to make a discovery, demonstrate a known fact, illustrate particular effects or test a hypothesis
- ◆ carrying out complex experiments in chemistry safely, recording systematic detailed observations and collecting data
- ◆ selecting information from a variety of sources and presenting detailed information appropriately, in a variety of forms
- ◆ processing and analysing chemical information and data (using calculations, significant figures and units, where appropriate)
- ◆ making reasoned predictions and generalisations from a range of evidence and/or information
- ◆ drawing valid conclusions and giving explanations supported by evidence and/or justification
- ◆ critically evaluating experimental procedures by identifying sources of uncertainty and suggesting and implementing improvements
- ◆ drawing on knowledge and understanding of chemistry to make accurate statements, describe complex information, provide detailed explanations and integrate knowledge
- ◆ communicating chemical findings and information fully and effectively
- ◆ analysing and evaluating scientific publications and media reports

Skills, knowledge and understanding for the course assessment

The following provides details of skills, knowledge and understanding sampled in the course assessment:

Inorganic chemistry

(a) Electromagnetic radiation and atomic spectra

Electromagnetic radiation can be described in terms of waves and characterised in terms of wavelength and/or frequency.

The relationship between these quantities is given by $c = f\lambda$.

The different types of radiation arranged in order of wavelength is known as the electromagnetic spectrum.

Wavelengths of visible light are normally expressed in nanometres (nm).

Electromagnetic radiation can be described as a wave (has a wavelength and frequency), and as a particle, and is said to have a dual nature.

When electromagnetic radiation is absorbed or emitted by matter it behaves like a stream of particles. These particles are known as photons.

A photon carries quantised energy proportional to the frequency of radiation.

When a photon is absorbed or emitted, energy is gained or lost by electrons within the substance.

The photons in high frequency radiation can transfer greater amounts of energy than photons in low frequency radiation.

The energy associated with a single photon is given by:

$$E = hf \text{ or } E = \frac{hc}{\lambda}$$

The energy associated with one mole of photons is given by:

$$E = Lhf \text{ or } E = \frac{Lhc}{\lambda}$$

Energy is often in units of kJ mol^{-1} .

When energy is transferred to atoms, electrons within the atoms may be promoted to higher energy levels.

An atom emits a photon of light energy when an excited electron moves from a higher energy level to a lower energy level.

The light energy emitted by an atom produces a spectrum that is made up of a series of lines at discrete (quantised) energy levels. This provides direct evidence for the existence of these energy levels.

Inorganic chemistry

Each element in a sample produces characteristic absorption and emission spectra. These spectra can be used to identify and quantify the element.

In absorption spectroscopy, electromagnetic radiation is directed at an atomised sample. Radiation is absorbed as electrons are promoted to higher energy levels.

An absorption spectrum is produced by measuring how the intensity of absorbed light varies with wavelength.

In emission spectroscopy, high temperatures are used to excite the electrons within atoms. As the electrons drop to lower energy levels, photons are emitted.

An emission spectrum of a sample is produced by measuring the intensity of light emitted at different wavelengths.

In atomic spectroscopy, the concentration of an element within a sample is related to the intensity of light emitted or absorbed.

(b) Atomic orbitals, electronic configurations and the periodic table

The discrete lines observed in atomic spectra can be explained if electrons, like photons, also display the properties of both particles and waves.

Electrons behave as standing (stationary) waves in an atom. These are waves that vibrate in time but do not move in space. There are different sizes and shapes of standing wave possible around the nucleus, known as orbitals. Orbitals can hold a maximum of two electrons.

The different shapes of orbitals are identified as s, p, d and f (knowledge of the shape of f orbitals is not required).

Electrons within atoms have fixed amounts of energy called quanta.

It is possible to describe any electron in an atom using four quantum numbers:

- ◆ the principal quantum number n indicates the main energy level for an electron and is related to the size of the orbital
- ◆ the angular momentum quantum number l determines the shape of the subshell and can have values from zero to $n-1$
- ◆ the magnetic quantum number m_l determines the orientation of the orbital and can have values between $-l$ and $+l$
- ◆ the spin magnetic quantum number m_s determines the direction of spin and can have values of $+\frac{1}{2}$ or $-\frac{1}{2}$

Inorganic chemistry

Electrons within atoms are arranged according to:

- ◆ the aufbau principle — electrons fill orbitals in order of increasing energy ('aufbau' means 'building up' in German)
- ◆ Hund's rule — when degenerate orbitals are available, electrons fill each singly, keeping their spins parallel before spin pairing starts
- ◆ the Pauli exclusion principle — no two electrons in one atom can have the same set of four quantum numbers, therefore, no orbital can hold more than two electrons and these two electrons must have opposite spins

In an isolated atom the orbitals within each subshell are degenerate.

The relative energies corresponding to each orbital can be represented diagrammatically using orbital box notation for the first four shells of a multi-electron atom.

Electronic configurations using spectroscopic notation and orbital box notation can be written for elements of atomic numbers 1 to 36.

The periodic table is subdivided into four blocks (s, p, d and f) corresponding to the outer electronic configurations of the elements within these blocks.

The variation in first, second and subsequent ionisation energies with increasing atomic number for the first 36 elements can be explained in terms of the relative stability of different subshell electronic configurations. This provides evidence for these electronic configurations. Anomalies in the trends of ionisation energies can be explained by considering the electronic configurations.

There is a special stability associated with half-filled and full subshells. The more stable the electronic configuration, the higher the ionisation energy.

VSEPR (valence shell electron pair repulsion) theory can be used to predict the shapes of molecules and polyatomic ions.

The number of electron pairs surrounding a central atom can be found by:

- ◆ taking the total number of valence (outer) electrons on the central atom and adding one for each atom attached
- ◆ adding an electron for every negative charge
- ◆ removing an electron for every positive charge
- ◆ dividing the total number of electrons by two to give the number of electron pairs

Electron pairs are negatively charged and repel each other. They are arranged to minimise repulsion and maximise separation.

Inorganic chemistry

The arrangement of electron pairs around a central atom is:

- ◆ linear for two electron pairs
- ◆ trigonal planar for three electron pairs
- ◆ tetrahedral for four electron pairs
- ◆ trigonal bipyramidal for five electron pairs
- ◆ octahedral for six electron pairs

Shapes of molecules or polyatomic ions are determined by the shapes adopted by the atoms present based on the arrangement of electron pairs. Electron dot diagrams can be used to show these arrangements.

Electron pair repulsions decrease in strength in the order:

non-bonding pair/non-bonding pair > non-bonding pair/bonding pair > bonding pair/bonding pair

(c) Transition metals

The d-block transition metals are metals with an incomplete d subshell in at least one of their ions.

The filling of the d orbitals follows the aufbau principle, with the exception of chromium and copper atoms.

These exceptions are due to the special stability associated with the d subshell being half-filled or completely filled.

When atoms from the first row of the transition elements form ions, it is the 4s electrons that are lost first rather than the 3d electrons.

An element is said to be in a particular oxidation state when it has a specific oxidation number.

The oxidation number can be determined by the following:

- ◆ uncombined elements have an oxidation number of 0
- ◆ ions containing single atoms have an oxidation number that is the same as the charge on the ion
- ◆ in most of its compounds, oxygen has an oxidation number of -2
- ◆ in most of its compounds, hydrogen has an oxidation number of +1
- ◆ the sum of all the oxidation numbers of all the atoms in a neutral compound must add up to zero
- ◆ the sum of all the oxidation numbers of all the atoms in a polyatomic ion must be equal to the charge on the ion

Inorganic chemistry

A transition metal can have different oxidation states in its compounds.

Compounds of the same transition metal in different oxidation states may have different colours.

Oxidation can be defined as an increase in oxidation number. Reduction can be considered as a decrease in oxidation number.

Changes in oxidation number of transition metal ions can be used to determine whether oxidation or reduction has occurred.

Compounds containing metals in high oxidation states are often oxidising agents, whereas compounds with metals in low oxidation states are often reducing agents.

Ligands may be negative ions or molecules with non-bonding pairs of electrons that they donate to the central metal atom or ion, forming dative covalent bonds.

Ligands can be classified as monodentate, bidentate, up to hexadentate.

It is possible to deduce the ligand classification from a formula or structure of the ligand or complex.

The total number of bonds from the ligands to the central transition metal is known as the coordination number.

Names and formulae can be written according to IUPAC rules for complexes containing:

- ◆ central metals that obey the normal IUPAC rules
- ◆ copper (cuprate) and iron (ferrate)
- ◆ ligands, including water, ammonia, halogens, cyanide, hydroxide, and oxalate

In a complex of a transition metal, the d orbitals are no longer degenerate.

Splitting of d orbitals to higher and lower energies occurs when the electrons present in approaching ligands cause the electrons in the orbitals lying along the axes to be repelled.

Ligands that cause a large difference in energy between subsets of d orbitals are strong field ligands. Weak field ligands cause a small energy difference.

Ligands can be placed in an order of their ability to split d orbitals. This is called the spectrochemical series.

Colours of many transition metal complexes can be explained in terms of d-d transitions.

Light is absorbed when electrons in a lower energy d orbital are promoted to a d orbital of higher energy.

Inorganic chemistry

If light of one colour is absorbed, then the complementary colour will be observed.

Electrons transition to higher energy levels when energy corresponding to the ultraviolet or visible regions of the electromagnetic spectrum is absorbed.

Transition metals and their compounds can act as catalysts.

Heterogeneous catalysts are in a different state to the reactants.

Heterogeneous catalysis can be explained in terms of the formation of activated complexes and the adsorption of reactive molecules onto active sites. The presence of unpaired d electrons or unfilled d orbitals is thought to allow activated complexes to form. This can provide reaction pathways with lower activation energies compared to the uncatalysed reaction.

Homogeneous catalysts are in the same state as the reactants.

Homogeneous catalysis can be explained in terms of changing oxidation states with the formation of intermediate complexes.

Physical chemistry

(a) Chemical equilibrium

A chemical reaction is in equilibrium when the composition of the reactants and products remains constant indefinitely.

The equilibrium constant (K) characterises the equilibrium composition of the reaction mixture.

For the general reaction $aA + bB \rightleftharpoons cC + dD$ the equilibrium expression is:

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$[A], [B], [C]$ and $[D]$ are the equilibrium concentrations of A, B, C and D and a, b, c and d are the stoichiometric coefficients in the balanced reaction equation.

The value of equilibrium constants can be calculated.

The value of an equilibrium constant indicates the position of equilibrium.

Equilibrium constants have no units.

The concentrations of pure solids and pure liquids at equilibrium are taken as constant and given a value of 1 in the equilibrium expression.

The numerical value of the equilibrium constant depends on the reaction temperature and is independent of concentration and/or pressure.

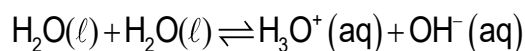
For endothermic reactions, a rise in temperature causes an increase in K and the yield of the product is increased.

For exothermic reactions, a rise in temperature causes a decrease in K and the yield of the product is decreased.

The presence of a catalyst does not affect the value of the equilibrium constant.

In water and aqueous solutions there is an equilibrium between the water molecules and hydronium (hydrogen) and hydroxide ions.

This ionisation of water can be represented by:



$\text{H}_3\text{O}^+(\text{aq})$ represents a hydronium ion, a hydrated proton. A shorthand representation of $\text{H}_3\text{O}^+(\text{aq})$ is $\text{H}^+(\text{aq})$.

Physical chemistry

Water is amphoteric (can react as an acid and a base).

The dissociation constant for the ionisation of water is known as the ionic product and is represented by K_w :

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

The value of the ionic product varies with temperature.

At 25°C the value of K_w is approximately 1×10^{-14} .

The relationship between pH and the hydrogen ion concentration is given by:

$$\text{pH} = -\log_{10} [\text{H}_3\text{O}^+] \text{ and } [\text{H}_3\text{O}^+] = 10^{-\text{pH}}$$

In water and aqueous solutions with a pH value of 7 the concentrations of $\text{H}_3\text{O}^+(\text{aq})$ and $\text{OH}^-(\text{aq})$ are both $10^{-7} \text{ mol l}^{-1}$ at 25°C.

If the concentration of $\text{H}_3\text{O}^+(\text{aq})$ or the concentration of $\text{OH}^-(\text{aq})$ is known, the concentration of the other ion can be calculated using K_w or by using $\text{pH} + \text{pOH} = 14$.

The Brønsted-Lowry definitions of acids and bases state that an acid is a proton donor and a base is a proton acceptor.

For every acid there is a conjugate base, formed by the loss of a proton.

For every base there is a conjugate acid, formed by the gain of a proton.

Strong acids and strong bases are completely dissociated into ions in aqueous solution.

Weak acids and weak bases are only partially dissociated into ions in aqueous solution.

Examples of strong acids include hydrochloric acid, sulfuric acid and nitric acid.

Ethanoic acid, carbonic acid and sulfurous acid are examples of weak acids.

Solutions of metal hydroxides are strong bases.

Ammonia and amines are examples of weak bases.

The weakly acidic nature of solutions of carboxylic acids, sulfur dioxide and carbon dioxide can be explained by reference to equations showing the equilibria.

The weakly alkaline nature of a solution of ammonia or amines can be explained by reference to an equation showing the equilibrium.

Physical chemistry

Equimolar solutions of weak and strong acids (or bases) have different pH values, conductivity, and reaction rates, but the stoichiometry of reactions are the same.

The acid dissociation constant is represented by K_a :

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

or by:

$$\text{p}K_a \text{ where } \text{p}K_a = -\log_{10} K_a$$

The approximate pH of a weak acid can be calculated using:

$$\text{pH} = \frac{1}{2}\text{p}K_a - \frac{1}{2}\log_{10} c .$$

A soluble salt of a strong acid and a strong base dissolves in water to produce a neutral solution.

A soluble salt of a weak acid and a strong base dissolves in water to produce an alkaline solution.

A soluble salt of a strong acid and a weak base dissolves in water to produce an acidic solution.

The name of the salt produced depends on the acid and base used.

Using the appropriate equilibria, the changes in concentrations of H_3O^+ and OH^- ions of salt solutions can be explained.

A buffer solution is one in which the pH remains approximately constant when small amounts of acid, base or water are added.

An acid buffer consists of a solution of a weak acid and one of its salts made from a strong base.

In an acid buffer solution the weak acid provides hydrogen ions when these are removed by the addition of a small amount of base. The salt of the weak acid provides the conjugate base, which can absorb excess hydrogen ions produced by the addition of a small amount of acid.

A basic buffer consists of a solution of a weak base and one of its salts.

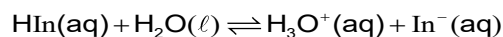
In a basic buffer solution the weak base removes excess hydrogen ions, and the conjugate acid provided by the salt supplies hydrogen ions when these are removed.

Physical chemistry

An approximate pH of an acid buffer solution can be calculated from its composition and from the acid dissociation constant:

$$\text{pH} = \text{p}K_a - \log_{10} \frac{[\text{acid}]}{[\text{salt}]}$$

Indicators are weak acids for which the dissociation can be represented as:



The acid indicator dissociation constant is represented as K_{In} and is given by the following expression:

$$K_{\text{In}} = \frac{[\text{H}_3\text{O}^+][\text{In}^-]}{[\text{HIn}]}$$

In aqueous solution the colour of an acid indicator is distinctly different from that of its conjugate base.

The colour of the indicator is determined by the ratio of $[\text{HIn}]$ to $[\text{In}^-]$.

The theoretical point at which colour change occurs is when $[\text{H}_3\text{O}^+] = K_{\text{In}}$.

The colour change is assumed to be distinguishable when $[\text{HIn}]$ and $[\text{In}^-]$ differ by a factor of 10.

The pH range over which a colour change occurs can be estimated by the expression:

$$\text{pH} = \text{p}K_{\text{In}} \pm 1$$

Suitable indicators can be selected from pH data, including titration curves.

Physical chemistry

(b) Reaction feasibility

The standard enthalpy of formation, ΔH_f° , is the enthalpy change when one mole of a substance is formed from its elements in their standard states.

The standard state of a substance is its most stable state at a pressure of 1 atmosphere and at a specified temperature, usually taken as 298 K.

The standard enthalpy of a reaction can be calculated from the standard enthalpies of formation of the reactants and products:

$$\Delta H^\circ = \sum \Delta H_f^\circ(\text{products}) - \sum \Delta H_f^\circ(\text{reactants})$$

The entropy (S) of a system is a measure of the degree of disorder of the system.

The greater the degree of disorder, the greater the entropy.

Solids have low disorder and gases have high disorder.

Entropy increases as temperature increases.

There is a rapid increase in entropy at the melting point of a substance and an even more rapid and larger change in entropy at the boiling point.

The second law of thermodynamics states that the total entropy of a reaction system and its surroundings always increases for a spontaneous process.

Heat energy released by the reaction system into the surroundings increases the entropy of the surroundings.

Heat energy absorbed by the reaction system from the surroundings decreases the entropy of the surroundings.

The third law of thermodynamics states that the entropy of a perfect crystal at 0 K is zero.

The standard entropy of a substance is the entropy value for the substance in its standard state.

The change in standard entropy for a reaction system can be calculated from the standard entropies of the reactants and products:

$$\Delta S^\circ = \sum S^\circ(\text{products}) - \sum S^\circ(\text{reactants})$$

The change in free energy for a reaction is related to the enthalpy and entropy changes:

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

Physical chemistry

If the change in free energy (ΔG°) between reactants and products is negative, a reaction may occur and the reaction is said to be feasible. A feasible reaction is one that tends towards the products rather than the reactants. This does not give any indication of the rate of the reaction.

The standard free energy change for a reaction can be calculated from the standard free energies of formation of the reactants and products using the relationship:

$$\Delta G^\circ = \sum \Delta G^\circ(\text{products}) - \sum \Delta G^\circ(\text{reactants})$$

The feasibility of a chemical reaction under standard conditions can be predicted from the calculated value of the change in standard free energy (ΔG°).

The temperatures at which a reaction may be feasible can be estimated by considering the range of values of T for which $\Delta G^\circ < 0$.

Under non-standard conditions any reaction is feasible if ΔG is negative.

At equilibrium, $\Delta G = 0$.

A reversible reaction will proceed spontaneously until the composition is reached where $\Delta G = 0$.

Physical chemistry

(c) Kinetics

The rate of a chemical reaction normally depends on the concentrations of the reactants.

Orders of reaction are used to relate the rate of a reaction to the reacting species.

If changing the concentration of a reactant A has no effect on the rate of the reaction, then the reaction is zero order with respect to A .

If doubling the concentration of a reactant A doubles the rate of the reaction, then the reaction is first order with respect to A . The rate can be expressed as:

rate = $k[A]$ where k is the rate constant and $[A]$ is the concentration of reactant A in mol l^{-1}

If doubling the concentration of a reactant A increases the rate of the reaction fourfold, then the reaction is second order with respect to A . The rate can be expressed as:

rate = $k[A]^2$

The order of a reaction with respect to any one reactant is the power to which the concentration of that reactant is raised in the rate equation.

The overall order of a reaction is the sum of the powers to which the concentrations of the reactants are raised in the rate equation.

The order of a reaction can only be determined from experimental data.

The rate equation and the rate constant, including units, can be determined from initial rate data for a series of reactions in which the initial concentrations of reactants are varied. These can be zero, first, second or third order.

Reactions usually occur by a series of steps called a reaction mechanism.

The rate of reaction is dependent on the slowest step, which is called the 'rate determining step'.

Experimentally determined rate equations can be used to determine possible reaction mechanisms.

Organic chemistry and instrumental analysis

(a) Molecular orbitals

VSEPR cannot explain the bonding in all compounds. Molecular orbital theory can provide an explanation for more complex molecules.

Molecular orbitals form when atomic orbitals combine. The number of molecular orbitals formed is equal to the number of atomic orbitals that combine. The combination of two atomic orbitals results in the formation of a bonding molecular orbital and an antibonding orbital. The bonding molecular orbital encompasses both nuclei. The attraction of the positively charged nuclei and the negatively charged electrons in the bonding molecular orbital is the basis of bonding between atoms. Each molecular orbital can hold a maximum of two electrons.

In a non-polar covalent bond, the bonding molecular orbital is symmetrical about the midpoint between two atoms. Polar covalent bonds result from bonding molecular orbitals that are asymmetric about the midpoint between two atoms. The atom with the greater value for electronegativity has the greater share of the bonding electrons. Ionic compounds are an extreme case of asymmetry, with the bonding molecular orbitals being almost entirely located around just one atom, resulting in the formation of ions.

Molecular orbitals that form by end-on overlap of atomic orbitals along the axis of the covalent bond are called sigma (σ) molecular orbitals or sigma bonds.

Molecular orbitals that form by side-on overlap of parallel atomic orbitals that lie perpendicular to the axis of the covalent bond are called pi (π) molecular orbitals or pi bonds.

The electronic configuration of an isolated carbon atom cannot explain the number of bonds formed by carbon atoms in molecules. The bonding and shape of molecules of carbon can be explained by hybridisation.

Hybridisation is the process of mixing atomic orbitals within an atom to generate a set of new atomic orbitals called hybrid orbitals. These hybrid orbitals are degenerate.

In alkanes, the 2s orbital and the three 2p orbitals of carbon hybridise to form four degenerate sp^3 hybrid orbitals. These adopt a tetrahedral arrangement. The sp^3 hybrid orbitals overlap end-on with other atomic orbitals to form σ bonds.

The bonding in alkenes can be described in terms of sp^2 hybridisation. The 2s orbital and two of the 2p orbitals hybridise to form three degenerate sp^2 hybrid orbitals. These adopt a trigonal planar arrangement. The hybrid sp^2 orbitals overlap end-on to form σ bonds. The remaining 2p orbital on each carbon atom of the double bond is unhybridised and lies perpendicular to the axis of the σ bond. The unhybridised p orbitals overlap side-on to form π bonds.

Organic chemistry and instrumental analysis

The bonding in benzene and other aromatic systems can be described in terms of sp^2 hybridisation. The six carbon atoms in benzene are arranged in a cyclic structure with σ bonds between the carbon atoms. The unhybridised p orbitals on each carbon atom overlap side-on to form a π molecular system, perpendicular to the plane of the σ bonds. This π molecular system extends across all six carbon atoms. The electrons in this system are delocalised.

The bonding in alkynes can be described in terms of sp hybridisation. The 2s orbital and one 2p orbital of carbon hybridise to form two degenerate hybrid orbitals. These adopt a linear arrangement. The hybrid sp orbitals overlap end-on to form σ bonds. The remaining two 2p orbitals on each carbon atom lie perpendicular to each other and to the axis of the σ bond. The unhybridised p orbitals overlap side-on to form two π bonds.

Molecular orbital theory can be used to explain why organic molecules are colourless or coloured. Electrons fill bonding molecular orbitals, leaving higher energy antibonding orbitals unfilled. The highest bonding molecular orbital containing electrons is called the highest occupied molecular orbital (HOMO). The lowest antibonding molecular orbital is called the lowest unoccupied molecular orbital (LUMO).

Absorption of electromagnetic energy can cause electrons to be promoted from HOMO to LUMO.

Most organic molecules appear colourless because the energy difference between HOMO and LUMO is relatively large. This results in absorption of light from the ultraviolet region of the spectrum.

Some organic molecules contain chromophores. A chromophore is a group of atoms within a molecule that is responsible for absorption of light in the visible region of the spectrum. Light can be absorbed when electrons in a chromophore are promoted from the HOMO to the LUMO.

Chromophores exist in molecules containing a conjugated system — a system of adjacent unhybridised p orbitals that overlap side-on to form a molecular orbital across a number of carbon atoms. Electrons within this conjugated system are delocalised. Molecules with alternating single and double bonds, and aromatic molecules have conjugated systems.

The more atoms in the conjugated system the smaller the energy gap between HOMO and LUMO. A lower frequency of light (longer wavelength, lower energy) is absorbed by the compound. When the wavelength of light absorbed is in the visible region, the compound will exhibit the complementary colour.

Organic chemistry and instrumental analysis

(b) Synthesis

When an organic reaction takes place, bonds in the reactant molecules are broken and bonds in the product molecules are made. The process of bond breaking is known as bond fission.

There are two types of bond fission, homolytic and heterolytic.

Homolytic fission:

- ◆ results in the formation of two neutral radicals
- ◆ occurs when each atom retains one electron from the σ covalent bond and the bond breaks evenly
- ◆ normally occurs when non-polar covalent bonds are broken

Reactions involving homolytic fission tend to result in the formation of very complex mixtures of products, making them unsuitable for organic synthesis.

Heterolytic fission:

- ◆ results in the formation of two oppositely charged ions
- ◆ occurs when one atom retains both electrons from the σ covalent bond and the bond breaks unevenly
- ◆ normally occurs when polar covalent bonds are broken

Reactions involving heterolytic fission tend to result in far fewer products than reactions involving homolytic fission, and so are better suited for organic synthesis.

The movement of electrons during bond fission and bond making can be represented using curly arrow notation where:

- ◆ a single-headed arrow indicates the movement of a single electron
- ◆ a double-headed arrow indicates the movement of an electron pair
- ◆ the tail of the arrow shows the source of the electron(s)
- ◆ the head of the arrow indicates the destination of the electron(s)
- ◆ two single-headed arrows starting at the middle of a covalent bond indicate homolytic bond fission is occurring
- ◆ a double-headed arrow starting at the middle of a covalent bond indicates heterolytic bond fission is occurring
- ◆ an arrow drawn with the head pointing to the space between two atoms indicates that a covalent bond will be formed between those two atoms

Organic chemistry and instrumental analysis

In reactions involving heterolytic bond fission, attacking groups are classified as nucleophiles or electrophiles.

Nucleophiles are:

- ◆ negatively charged ions or neutral molecules that are electron rich, such as, Cl^- , Br^- , OH^- , CN^- , NH_3 and H_2O
- ◆ attracted towards atoms bearing a partial (δ^+) or full positive charge
- ◆ capable of donating an electron pair to form a new covalent bond

Electrophiles are:

- ◆ positively charged ions or neutral molecules that are electron deficient, such as H^+ , NO_2^+ and SO_3
- ◆ attracted towards atoms bearing a partial (δ^-) or full negative charge
- ◆ capable of accepting an electron pair to form a new covalent bond

The following reaction types can be identified from a chemical equation:

- ◆ substitution
- ◆ addition
- ◆ elimination
- ◆ condensation
- ◆ hydrolysis
- ◆ oxidation
- ◆ reduction
- ◆ neutralisation

Synthetic routes can be devised, with no more than three steps, from a given reactant to a final product.

The possible reactions of a particular molecule can be deduced by looking at the structural formula.

The structure of any molecule can be drawn as a full, shortened or skeletal structural formula.

In a skeletal structural formula, neither the carbon atoms, nor any hydrogens attached to the carbon atoms, are shown. The presence of a carbon atom is implied by a 'kink' in the carbon backbone, and at the end of a line.

Given a full or shortened structural formula for a compound, the skeletal structural formula can be drawn.

Organic chemistry and instrumental analysis

Given a skeletal structural formula for a compound, the full or shortened structural formula can be drawn.

Molecular formulae can be written from a full, shortened or skeletal structural formula.

Straight and branched chain alkanes; alkenes; alcohols; carboxylic acids; aldehydes and ketones; haloalkanes; and ethers can be systematically named, indicating the position of the functional group where appropriate, from structural formulae containing no more than eight carbon atoms in their longest chain. Straight chain esters can be systematically named from the names of their parent alcohol and carboxylic acid or their structural formula.

Molecular formulae can be written and structural formulae drawn from systematic names of straight and branched chain alkanes; alkenes; alcohols; carboxylic acids; aldehydes and ketones; haloalkanes; and ethers containing no more than eight carbon atoms in their longest chain. Molecular formulae can be written and structural formulae drawn for esters from the systematic name or the structural formulae of their parent alcohol and carboxylic acid.

Haloalkanes (alkyl halides) are substituted alkanes in which one or more of the hydrogen atoms is replaced with a halogen atom.

Monohaloalkanes:

- ◆ contain only one halogen atom
- ◆ can be classified as primary, secondary or tertiary according to the number of alkyl groups attached to the carbon atom containing the halogen atom
- ◆ take part in elimination reactions to form alkenes using a strong base, such as potassium or sodium hydroxide in ethanol
- ◆ take part in nucleophilic substitution reactions with:
 - aqueous alkalis to form alcohols
 - alcoholic alkoxides to form ethers
 - ethanolic cyanide to form nitriles (chain length increased by one carbon atom) that can be hydrolysed to carboxylic acids

A monohaloalkane can take part in nucleophilic substitution reactions by one of two different mechanisms.

S_N1 is a nucleophilic substitution reaction with one species in the rate determining step and occurs in a minimum of two steps via a trigonal planar carbocation intermediate.

S_N2 is a nucleophilic substitution reaction with two species in the rate determining step and occurs in a single step via a single five-centred, trigonal bipyramidal transition state.

Organic chemistry and instrumental analysis

The reaction mechanisms for S_N1 and S_N2 reactions can be represented using curly arrows. Steric hindrance and the inductive stabilisation of the carbocation intermediate can be used to explain which mechanism will be preferred for a given haloalkane.

Alcohols are substituted alkanes in which one or more of the hydrogen atoms is replaced with a hydroxyl functional group, $-OH$ group.

Alcohols can be prepared from:

- ◆ haloalkanes by substitution
- ◆ alkenes by acid-catalysed hydration (addition)
- ◆ aldehydes and ketones by reduction using a reducing agent such as lithium aluminium hydride

Reactions of alcohols include:

- ◆ dehydration to form alkenes using aluminium oxide, concentrated sulfuric acid or concentrated phosphoric acid
- ◆ oxidation of primary alcohols to form aldehydes and then carboxylic acids and secondary alcohols to form ketones, using acidified permanganate, acidified dichromate or hot copper(II) oxide
- ◆ formation of alcoholic alkoxides by reaction with some reactive metals such as potassium or sodium, which can then be reacted with monohaloalkanes to form ethers
- ◆ formation of esters by reaction with carboxylic acids using concentrated sulfuric acid or concentrated phosphoric acid as a catalyst
- ◆ formation of esters by reaction with acid chlorides (see diagram below) — this gives a faster reaction than reaction with carboxylic acids, and no catalyst is needed



Hydroxyl groups make alcohols polar, which gives rise to hydrogen bonding. Hydrogen bonding can be used to explain the properties of alcohols including boiling points, melting points, viscosity and solubility or miscibility in water.

Ethers can be regarded as substituted alkanes in which a hydrogen atom is replaced with an alkoxy functional group, $-OR$, and have the general structure $R' - O - R''$, where R' and R'' are alkyl groups.

Ethers are named as substituted alkanes. The alkoxy group is named by adding the ending 'oxy' to the alkyl substituent, and this prefixes the name of the longest carbon chain.

Organic chemistry and instrumental analysis

Ethers can be prepared in a nucleophilic substitution reaction by reacting a monohaloalkane with an alkoxide.

Due to the lack of hydrogen bonding between ether molecules, they have lower boiling points than the corresponding isomeric alcohols.

Methoxymethane and methoxyethane are soluble in water. Larger ethers are insoluble in water due to their increased molecular size.

Ethers are commonly used as solvents since they are relatively inert chemically and will dissolve many organic compounds.

Alkenes can be prepared by:

- ◆ dehydration of alcohols using aluminium oxide, concentrated sulfuric acid or concentrated phosphoric acid
- ◆ base-induced elimination of hydrogen halides from monohaloalkanes

Alkenes take part in electrophilic addition reactions with:

- ◆ hydrogen to form alkanes in the presence of a catalyst
- ◆ halogens to form dihaloalkanes
- ◆ hydrogen halides to form monohaloalkanes
- ◆ water using an acid catalyst to form alcohols

Markovnikov's rule states that when a hydrogen halide or water is added to an unsymmetrical alkene, the hydrogen atom becomes attached to the carbon with the most hydrogen atoms attached to it already. Markovnikov's rule can be used to predict major and minor products formed during the reaction of a hydrogen halide or water with alkenes.

The reaction mechanisms for the addition of a hydrogen halide and the acid-catalysed addition of water can be represented using curly arrows and showing the intermediate carbocation. The inductive stabilisation of intermediate carbocations formed during these reactions can be used to explain the products formed.

The reaction mechanism for the addition of a halogen can be represented using curly arrows and showing the cyclic ion intermediate.

Carboxylic acids can be prepared by:

- ◆ oxidising primary alcohols using acidified permanganate, acidified dichromate and hot copper(II) oxide
- ◆ oxidising aldehydes using acidified permanganate, acidified dichromate, Fehling's solution and Tollens' reagent
- ◆ hydrolysing nitriles, esters or amides

Organic chemistry and instrumental analysis

Reactions of carboxylic acids include:

- ◆ formation of salts by reactions with metals or bases
- ◆ condensation reactions with alcohols to form esters in the presence of concentrated sulfuric or concentrated phosphoric acid
- ◆ reaction with amines to form alkylammonium salts that form amides when heated
- ◆ reduction with lithium aluminium hydride to form primary alcohols

Amines are organic derivatives of ammonia in which one or more hydrogen atoms of ammonia has been replaced by an alkyl group.

Amines can be classified as primary, secondary or tertiary according to the number of alkyl groups attached to the nitrogen atom.

Amines react with acids to form salts.

Primary and secondary amines, but not tertiary amines, display hydrogen bonding. As a result, primary and secondary amines have higher boiling points than isomeric tertiary amines.

Primary, secondary and tertiary amine molecules can hydrogen-bond with water molecules, thus explaining the appreciable solubility of the shorter chain length amines in water.

Amines like ammonia are weak bases and dissociate to a slight extent in aqueous solution. The nitrogen atom has a lone pair of electrons which can accept a proton from water, producing hydroxide ions.

Benzene (C_6H_6) is the simplest member of the class of aromatic hydrocarbons.

The benzene ring has a distinctive structural formula. The stability of the benzene ring is due to the delocalisation of electrons in the conjugated system. The presence of delocalised electrons explains why the benzene ring does not take part in addition reactions.

Bonding in benzene can be described in terms of sp^2 hybridisation, sigma and pi bonds, and electron delocalisation.

A benzene ring in which one hydrogen atom has been substituted by another group is known as the phenyl group. The phenyl group has the formula $-C_6H_5$.

Organic chemistry and instrumental analysis

Benzene rings can take part in electrophilic substitution reactions. Reactions at benzene rings include:

- ◆ halogenation by reaction of a halogen using aluminium chloride or iron(III) chloride for chlorination and aluminium bromide or iron(III) bromide for bromination
- ◆ alkylation by reaction of a haloalkane using aluminium chloride
- ◆ nitration using concentrated sulfuric acid and concentrated nitric acid
- ◆ sulfonation using concentrated sulfuric acid

Organic chemistry and instrumental analysis

(c) Stereo chemistry

Molecules that have the same molecular formula but different structural formulae are called isomers.

Structural isomers occur when the atoms are bonded together in a different order in each isomer.

Stereoisomers occur when the order of the bonding in the atoms is the same but the spatial arrangement of the atoms is different in each isomer. There are two types of stereoisomer, geometric and optical.

Geometric isomers:

- ◆ can occur when there is restricted rotation around a carbon-carbon double bond or a carbon-carbon single bond in a cyclic compound
- ◆ must have two different groups attached to each of the carbon atoms that make up the bond with restricted rotation
- ◆ can be labelled cis or trans according to whether the substituent groups are on the same side (cis) or on different sides (trans) of the bond with restricted rotation
- ◆ have differences in physical properties, such as melting point and boiling point
- ◆ can have differences in chemical properties

Optical isomers:

- ◆ occur in compounds in which four different groups are arranged tetrahedrally around a central carbon atom (chiral carbon or chiral centre)
- ◆ are asymmetric
- ◆ are non-superimposable mirror images of each other
- ◆ can be described as enantiomers
- ◆ have identical physical properties, except for their effect on plane-polarised light
- ◆ have identical chemical properties, except when in a chiral environment such as that found in biological systems (only one optical isomer is usually present)
- ◆ rotate plane-polarised light by the same amount but in opposite directions and so are optically active
- ◆ when mixed in equal amounts are optically inactive because the rotational effect of the plane-polarised light cancels out — this is called a racemic mixture

Organic chemistry and instrumental analysis

(d) Experimental determination of structure

In organic chemistry a number of experimental techniques are carried out to verify the chemical structure of a substance.

Elemental microanalysis is used to determine the masses of C, H, O, S and N in a sample of an organic compound in order to determine its empirical formula.

An empirical formula shows the simplest ratio of the elements in a molecule.

Elemental microanalysis can be determined from:

- ◆ combustion product masses
- ◆ percentage product by mass

Mass spectrometry can be used to determine the accurate gram formula mass (GFM) and structural features of an organic compound.

In mass spectrometry, a small sample of an organic compound is bombarded by high-energy electrons. This removes electrons from the organic molecule generating positively charged molecular ions known as parent ions. These molecular ions then break into smaller positively charged ion fragments. A mass spectrum is obtained showing a plot of the relative abundance of the ions detected against the mass-to-charge (m/z) ratio.

The mass-to-charge ratio of the parent ion can be used to determine the GFM of the molecular ion, and so a molecular formula can be determined using the empirical formula.

The fragmentation data can be interpreted to gain structural information.

Infrared spectroscopy is used to identify certain functional groups in an organic compound.

When infrared radiation is absorbed by organic compounds, bonds within the molecule vibrate (stretch and bend). The wavelengths of infrared radiation that are absorbed depend on the type of atoms that make up the bond and the strength of the bond.

In infrared spectroscopy, infrared radiation is passed through a sample of the organic compound and then into a detector that measures the intensity of the transmitted radiation at different wavelengths. The absorbance of infrared radiation is measured in wavenumbers, the reciprocal of wavelength, in units of cm^{-1} .

Characteristic absorptions by particular vibrations are given in the data booklet.

Organic chemistry and instrumental analysis

Proton nuclear magnetic resonance spectroscopy (proton NMR or ^1H NMR) can give information about the different chemical environments of hydrogen atoms (protons or ^1H) in an organic molecule, and about how many hydrogen atoms there are in each of these environments.

^1H nuclei behave like tiny magnets and in a strong magnetic field some align with the field (lower energy), whilst the rest align against it (higher energy). Absorption of radiation in the radio frequency region of the electromagnetic spectrum causes the ^1H nuclei to 'flip' from the lower to the higher energy alignment. As they fall back from the higher to the lower energy alignment the emitted radiation is detected and plotted on a spectrum.

In a ^1H NMR spectrum the chemical shift, δ , (peak position) is related to the environment of the ^1H atom and is measured in parts per million (ppm).

Chemical shift values for ^1H in different chemical environments are given in the data booklet.

The area under the peak is related to the number of ^1H atoms in that environment and is often given by an integration curve on a spectrum. The height of an integration curve is proportional to the number of ^1H atoms in that environment, and so a ratio of ^1H atoms in each environment can be determined.

The standard reference substance used in ^1H NMR spectroscopy is tetramethylsilane (TMS), which is assigned a chemical shift value equal to zero.

^1H NMR spectra can be obtained using low-resolution or high-resolution NMR.

High-resolution ^1H NMR uses higher radio frequencies than those used in low-resolution ^1H NMR and provides more detailed spectra.

In a high-resolution ^1H NMR an interaction with ^1H atoms on neighbouring carbon atoms can result in the splitting of peaks into multiplets. The number of ^1H atoms on neighbouring carbon atoms will determine the number of peaks within a multiplet and can be determined using the $n+1$ rule, where n is the number of ^1H atoms on the neighbouring carbon atom.

Low- and high-resolution ^1H NMR spectra can be analysed, and low-resolution ^1H NMR spectra can be sketched for any given compound.

Organic chemistry and instrumental analysis

(e) Pharmaceutical chemistry

Drugs are substances that alter the biochemical processes in the body.

Drugs that have beneficial effects are used in medicines.

A medicine usually contains the drug plus other ingredients such as fillers to add bulk or sweeteners to improve the taste.

Drugs generally work by binding to specific protein molecules. These protein molecules can be found on the surface of a cell (receptor) or can be specific enzyme molecules within a cell.

Drugs that act on receptors can be classified as agonists or antagonists.

- ◆ An agonist mimics the natural compound and binds to the receptor molecules to produce a response similar to the natural active compound.
- ◆ An antagonist prevents the natural compound from binding to the receptor, and so blocks the natural response from occurring.

Many drugs that act on enzymes are classified as enzyme inhibitors and act by binding to the active site of the enzyme and blocking the reaction normally catalysed there.

The overall shape and size of a drug is such that it interacts with a receptor binding site or to the active site of an enzyme. The types of interactions formed can include Van der Waals forces and/or ionic bonds.

The structural fragment of a drug molecule that allows it to form interactions with a receptor binding site or to an enzyme active site normally consists of different functional groups correctly orientated with respect to each other.

By comparing the structures of drugs that have similar effects on the body, the structural fragment that is involved in the drug action can be identified.

Researching chemistry

(a) Common chemical apparatus

Candidates must be familiar with the use(s) of the following types of apparatus:

- ◆ conical flask
- ◆ digital balance
- ◆ pipette with safety filler
- ◆ burette
- ◆ volumetric (standard) flask
- ◆ distillation (round-bottomed) flask
- ◆ condenser
- ◆ thermometer
- ◆ Büchner or Hirsch or sintered glass funnel
- ◆ glassware with ground glass joints ('Quickfit' or similar)
- ◆ thin-layer chromatography apparatus
- ◆ colorimeter
- ◆ melting point
- ◆ separating funnel

(b) Skills involved in experimental work

Candidates must be able to:

- ◆ tabulate data using appropriate headings and units of measurement
- ◆ represent data as a scatter graph with suitable scales and labels
- ◆ sketch a line of best fit (straight or curved) to represent the trend observed in the data
- ◆ calculate average (mean) values
- ◆ identify and eliminate rogue points
- ◆ qualitatively appreciate the relative accuracy of apparatus used to measure the volume of liquids
- ◆ comment on the reproducibility of results where measurements have been repeated
- ◆ carry out quantitative stoichiometric calculations
- ◆ interpret spectral data
- ◆ appropriately use a positive control, for example a known substance, to validate a technique or procedure

Researching chemistry

(c) Stoichiometric calculations

Stoichiometry is the study of mole relationships involved in chemical reactions.

Chemical equations, using formulae and state symbols, can be written and balanced to show the mole ratio(s) of reactants and products, including multi-step reactions.

The mass of a mole of any substance, in grams (g), is equal to the gram formula mass (GFM) and can be calculated using relative atomic masses.

Calculations can be performed using the relationship between the mass and the number of moles of a substance.

For solutions, the mass of solute (grams or g), the number of moles of solute (moles or mol), the volume of solution (litres or l), or the concentration of the solution (moles per litre or mol l^{-1}), can be calculated from data provided.

Percentage by mass is the mass of solute made up to 100 cm^3 of solution.

Percentage by volume is the number of cm^3 of solute made up to 100 cm^3 of solution.

The unit ppm stands for parts per million and refers to 1 mg per kg or 1 mg per litre.

Calculations can be performed using data, including:

- ◆ GFM
- ◆ masses
- ◆ number of moles
- ◆ concentrations and volumes of solutions
- ◆ volumes of gases
- ◆ reactant excess
- ◆ theoretical and percentage yield
- ◆ empirical formulae

Theoretical yields can be calculated and compared with actual yields, leading to determining the percentage yield. The percentage yield is reduced by:

- ◆ mass transfer or mechanical losses
- ◆ purification of product
- ◆ side reactions
- ◆ equilibrium position

Candidates must be able to carry out stoichiometric calculations for all of the skills and techniques in the course where appropriate.

Researching chemistry

(d) Gravimetric analysis

Candidates must be familiar with the technique of gravimetric analysis, including use of:

- ◆ an accurate electronic balance, including the tare function
- ◆ a weighing boat
- ◆ weighing by difference
- ◆ the term 'weighing accurately approximately'
- ◆ heating to constant mass:
 - heating a substance
 - allowing to cool in a desiccator to prevent absorption of water
 - weighing
 - repeating the steps of heating, cooling and weighing until no further changes in mass are observed

Gravimetric analysis is used to determine the mass of an element or compound in a substance.

The substance is converted into another substance of known chemical composition, which can be readily isolated and purified.

The conversion can occur either through precipitation or volatilisation.

In precipitation conversion the substance undergoes a precipitation reaction. The precipitate is separated from the filtrate and the filtrate tested to ensure the reaction has gone to completion. The precipitate is washed, dried to constant mass and then weighed.

In volatilisation conversion the substance is heated and any volatile products (often water) are evaporated. The substance is heated to constant mass and the final mass recorded.

(e) Volumetric analysis

Candidates must be familiar with use of the technique of volumetric analysis, including:

- ◆ preparing a standard solution
- ◆ accurate dilution
- ◆ standardising solutions to determine accurate concentration
- ◆ titrating to obtain concordancy using burettes, pipettes and volumetric flasks
- ◆ choosing an appropriate indicator

Researching chemistry

A solution of accurately known concentration is known as a standard solution.

A standard solution can be prepared by:

- ◆ weighing a primary standard accurately
- ◆ dissolving in a small volume of solvent (usually deionised or distilled water) in a beaker
- ◆ transferring the solution and rinsings into a volumetric flask
- ◆ making up to the graduation mark with solvent
- ◆ stoppering and inverting

Standard solutions can also be prepared by accurate dilution by pipetting an appropriate volume of a standard solution into a volumetric flask, making up to the graduation mark with solvent, stoppering and inverting.

A primary standard must:

- ◆ be available in a high state of purity
- ◆ be stable when solid and in solution
- ◆ be soluble
- ◆ have a reasonably high GFM

Examples of primary standards include:

- ◆ sodium carbonate, Na_2CO_3
- ◆ hydrated oxalic acid, $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$
- ◆ potassium hydrogen phthalate, $\text{KH}(\text{C}_8\text{H}_4\text{O}_4)$
- ◆ silver nitrate, AgNO_3
- ◆ potassium iodate, KIO_3
- ◆ potassium dichromate, $\text{K}_2\text{Cr}_2\text{O}_7$

Sodium hydroxide is not a primary standard as it has a relatively low GFM, is unstable as a solid (absorbs moisture) and unstable as a solution. Sodium hydroxide solution must be standardised before being used in volumetric analysis.

Researching chemistry

Candidates must be familiar with use of the following types of volumetric analysis:

- ◆ acid-base titrations
- ◆ redox titrations based on reactions between oxidising and reducing agents
- ◆ complexometric titrations based on reactions in which complexes are formed — EDTA is an important complexometric reagent and can be used to determine the concentration of metal ions in solution
- ◆ back titrations used to find the number of moles of a substance by reacting it with an excess volume of a reactant of known concentration. The resulting mixture is then titrated to work out the number of moles of the reactant in excess. From the initial number of moles of that reactant, the number of moles used in the reaction can be determined. The initial number of moles of the substance being analysed can then be calculated. A back titration is useful when trying to work out the quantity of substance in a solid with a low solubility.

(f) Practical skills and techniques

Candidates must be familiar with use of the technique of colorimetry, including:

- ◆ preparing a series of standard solutions of appropriate concentration
- ◆ choosing an appropriate colour or wavelength of filter complementary to the colour of the species being tested
- ◆ using a blank
- ◆ preparing a calibration graph

Colorimetry uses the relationship between colour intensity of a solution and the concentration of the coloured species present.

A colorimeter or a spectrophotometer is used to measure the absorbance of light of a series of standard solutions, and this data is used to plot a calibration graph.

The concentration of the solution being tested is determined from its absorbance and by referring to the calibration curve.

The concentration of coloured species in the solution being tested must lie in the straight line section of the calibration graph.

Candidates must be familiar with use of the technique of distillation. Distillation is used for identification and purification of organic compounds.

The boiling point of a compound, determined by distillation, is one of the physical properties that can be used to confirm its identity.

Distillation can be used to purify a compound by separating it from less volatile substances in the mixture.

Researching chemistry

Candidates must be familiar with use of the technique of heating under reflux. Heating under reflux allows heat energy to be applied to a chemical reaction mixture over an extended period of time without volatile substances escaping.

When carrying out heating under reflux, the reaction mixture is placed in a round-bottomed flask with anti-bumping granules and the flask is fitted with a condenser. The flask is then heated using an appropriate source of heat.

Candidates must be familiar with use of the technique of vacuum filtration. Vacuum filtration involves carrying out a filtration under reduced pressure and provides a faster means of separating a precipitate from a filtrate. A Büchner, Hirsch or sintered glass funnel can be used during vacuum filtration.

Candidates must be familiar with use of the technique of recrystallisation to purify an impure solid involving:

- ◆ dissolving an impure solid gently in a minimum volume of a hot solvent
- ◆ hot filtration of the resulting mixture to remove any insoluble impurities
- ◆ cooling the filtrate slowly to allow crystals of the pure compound to form, leaving soluble impurities dissolved in the solvent
- ◆ filtering, washing and drying the pure crystals

The solvent for recrystallisation is chosen so that the compound being purified is completely soluble at high temperatures and only sparingly soluble at lower temperatures.

Candidates must be familiar with use of the technique of solvent extraction. Solvent extraction involves isolating a solute from a liquid mixture or solution by extraction using an immiscible solvent in which the solute is soluble.

When carrying out a solvent extraction, the two immiscible solvents form two layers in the separating funnel. The solute dissolves in both solvents and an equilibrium establishes between the two layers. The ratio of solute dissolved in each layer is determined by the equilibrium constant, K . The lower layer is run off into a container and the upper layer is poured into a second container. This process is repeated to maximise the quantity of solute extracted.

Researching chemistry

The quantity of solute extracted is greater if a number of extractions using smaller volumes of solvent are carried out rather than a single extraction using a large volume of solvent.

The solvent used should be:

- ◆ immiscible with the liquid mixture or solution (usually water)
- ◆ one in which the solute is more soluble than in the liquid mixture or solution (usually water)
- ◆ volatile to allow the solute to be obtained by evaporation of the solvent
- ◆ unreactive with the solute

Candidates must be familiar with use of the techniques of melting point and mixed melting point determination. The melting point of a substance is the temperature range over which the solid first starts to melt, to when all of the solid has melted.

The identity of a pure compound can be confirmed by melting point analysis and a comparison of the experimentally determined melting point with a literature or known melting point value.

Determination of the melting point of a compound can give an indication of the purity of a compound. The presence of impurities in the compound lowers the melting point and broadens its melting temperature range due to the disruption in intermolecular bonding in the crystal lattice.

Determination of a mixed melting point involves mixing a small quantity of the product with some of the pure compound and determining the melting point. The melting point value and the range of the melting temperature can be used to determine if the product and the pure compound are the same substance.

Candidates must be familiar with use of the technique of thin-layer chromatography. Chromatography is a technique used to separate the components present within a mixture. Chromatography separates substances by making use of differences in their polarity or molecular size.

Thin-layer chromatography (TLC) uses a fine film of silica or aluminium oxide spread over glass, aluminium foil or plastic. A small sample of the mixture being tested is spotted onto the base (pencil) line of the chromatogram. A solvent dissolves the compounds in the spot and carries the compounds up the chromatogram. How far the compounds are carried depends on how soluble the compounds are in the chosen solvent and how well they adhere to the plate. A developing agent or ultraviolet light is normally required to visualise the spots on the chromatogram.

Researching chemistry

R_f values can be calculated:

$$R_f = \frac{\text{distance travelled by the sample}}{\text{distance travelled by the solvent}}$$

Under the same conditions (temperature, solvent, and saturation levels) a compound always has the same R_f value (within experimental error).

The identity of a compound can be confirmed by:

- ◆ comparing the experimentally determined R_f values with a literature or known value determined under the same conditions
- ◆ making a direct comparison on a TLC plate between the compound being tested and the pure substance — a co-spot could be used

TLC is used to assess the purity of substances. A pure substance, when spotted and developed on a TLC plate, should appear as a single spot (some impurities may not be visible by TLC analysis). The presence of more than one spot shows that impurities are present.

Skills, knowledge and understanding included in the course are appropriate to the SCQF level of the course. The SCQF level descriptors give further information on characteristics and expected performance at each SCQF level, and are available on the SCQF website.

Skills for learning, skills for life and skills for work

This course helps candidates to develop broad, generic skills. These skills are based on the [Skills for Learning, Skills for Life and Skills for Work](#) framework and draw from the following main skills areas:

1 Literacy

1.1 Reading

1.2 Writing

2 Numeracy

2.1 Number processes

2.2 Money, time and measurement

2.3 Information handling

5 Thinking skills

5.3 Applying

5.4 Analysing and evaluating

5.5 Creating

Teachers and/or lecturers must build these skills into the course at an appropriate level, where there are suitable opportunities.

Learning for Sustainability

Qualifications Scotland is committed to identifying opportunities to develop the values, skills, knowledge and understanding of Learning for Sustainability within all National Courses.

Course assessment

Course assessment is based on the information in this course specification.

The course assessment meets the purposes and aims of the course by addressing:

- ◆ breadth — drawing on knowledge and skills from across the course
- ◆ challenge — requiring greater depth or extension of knowledge and/or skills
- ◆ application — requiring application of knowledge and/or skills in practical or theoretical contexts as appropriate

This enables candidates to apply:

- ◆ breadth and depth of skills, knowledge and understanding from across the course to answer questions in chemistry
- ◆ skills of scientific inquiry, using related knowledge, to carry out a meaningful and appropriately challenging project in chemistry and communicate findings

The course assessment has two components, a question paper and a project. The relationship between these two components is complementary, to ensure appropriate coverage of the knowledge and skills of the course.

Course assessment structure: question paper

Question paper

100 marks

The question paper has a total mark allocation of 100 marks. This is 75% of the overall marks for the course assessment.

The question paper has two sections:

- ◆ Section 1 is an objective test and has 25 marks.
- ◆ Section 2 contains restricted-response questions and extended-response questions and has 75 marks.

A data booklet containing relevant data and formulae is provided.

The majority of the marks are awarded for applying knowledge and understanding. The other marks are awarded for applying skills of scientific inquiry, scientific analytical thinking and problem solving.

The question paper gives candidates an opportunity to demonstrate the following skills, knowledge and understanding:

- ◆ making accurate statements
- ◆ describing information, providing explanations and integrating knowledge

- ◆ applying knowledge of chemistry to new situations, interpreting information and solving problems
- ◆ planning and designing chemical experiments/investigations, including safety measures to make a discovery, demonstrate a known fact, illustrate particular effects or test a hypothesis
- ◆ selecting information from a variety of sources
- ◆ presenting information appropriately in a variety of forms
- ◆ processing information and data (using calculations, significant figures and units, where appropriate)
- ◆ making predictions and generalisations based on evidence and/or information
- ◆ drawing valid conclusions and giving explanations supported by evidence and/or justification
- ◆ evaluating experiments and suggesting improvements

Setting, conducting and marking the question paper

The question paper is set and marked by Qualifications Scotland, and conducted in centres under conditions specified for external examinations by Qualifications Scotland. Candidates have 2 hours and 30 minutes to complete the question paper.

Specimen question papers for Advanced Higher courses are published on Qualifications Scotland's website. These illustrate the standard, structure and requirements of the question papers. The specimen papers also include marking instructions.

Course assessment structure: project

Project

25 marks

The project has 25 marks. This is scaled by Qualifications Scotland to represent 25% of the overall marks for the course assessment.

The project allows candidates to carry out an in-depth investigation of a chemistry topic and produce a project report. Candidates are required to **individually** plan and carry out a chemistry investigation.

Candidates should keep a record of their work (lab book) as this will form the basis of their project report. This record should include details of their research, experiments and recorded data.

The project assesses the application of skills of scientific inquiry and related chemistry knowledge and understanding. It gives candidates an opportunity to demonstrate the following skills, knowledge and understanding:

- ◆ extending and applying knowledge of chemistry to new situations, interpreting and analysing information to solve more complex problems

- ◆ planning, designing and safely carrying out chemical experiments/investigations, including risk assessments to make a discovery, demonstrate a known fact, illustrate particular effects or test a hypothesis
- ◆ recording systematic detailed observations and collecting data
- ◆ selecting information from a variety of sources
- ◆ presenting detailed information appropriately in a variety of forms
- ◆ processing and analysing chemical information and data (using calculations, significant figures and units, where appropriate)
- ◆ making reasoned predictions and generalisations from a range of evidence and/or information
- ◆ drawing valid conclusions and giving explanations supported by evidence and/or justification
- ◆ critically evaluating experimental procedures by identifying sources of uncertainty and suggesting and implementing improvements
- ◆ drawing on knowledge and understanding of chemistry to make accurate statements, describe complex information, provide detailed explanations and integrate knowledge
- ◆ communicating chemical findings and information fully and effectively
- ◆ analysing and evaluating scientific publications and media reports

Project overview

Candidates carry out an in-depth investigation of a chemistry topic. Candidates choose their topic and **individually** investigate/research its underlying chemistry. Candidates must discuss potential topics with their teacher or lecturer to ensure that they do not waste time researching unsuitable topics. This is an open-ended task that may involve candidates carrying out a significant part of the work without close supervision.

Throughout the project candidates work autonomously, making independent and rational decisions based on evidence and interpretation of scientific information, which involves analysing and evaluating results. Through this, candidates further develop and enhance their scientific literacy skills.

The project offers challenge by requiring candidates to apply skills, knowledge and understanding in a context that is one or more of the following:

- ◆ unfamiliar
- ◆ familiar but investigated in greater depth
- ◆ integrating a number of familiar contexts

Candidates will produce a project report that has a logical structure.

Refer to the *Advanced Higher Chemistry Project Assessment Task* for detailed advice on the content of the project report.

Setting, conducting and marking the project

Setting

The project is set:

- ◆ by centres within Qualifications Scotland guidelines
- ◆ at a time appropriate to the candidate's needs
- ◆ within teaching and learning and includes experimental work at a level appropriate to Advanced Higher

Conducting

The project is conducted:

- ◆ under some supervision and control
- ◆ in time to meet a submission date set by Qualifications Scotland
- ◆ independently by the candidate

Marking

The project has 25 marks.

The table below gives the mark allocation for each assessment category of the project report.

Section	Expected response	Mark allocation
Abstract	A brief abstract (summary) stating the overall aim and conclusion of the project	1
Underlying chemistry	A description of the underlying chemistry that: <ul style="list-style-type: none">◆ is relevant to the project◆ demonstrates an understanding of the underlying chemistry relevant to the aim◆ demonstrates an understanding of the underlying chemistry of the techniques and procedures used◆ uses chemical terms/ideas accurately	3
Data collection and handling	Procedure(s) clearly described in the past tense and use the impersonal voice	2
Data collection and handling	Statement of hazard(s) for each chemical used and/or produced and additional safety measure(s)	1
Data collection and handling	Data collected using methods of appropriate complexity, including one procedure and at least one of the following: <ul style="list-style-type: none">◆ a second procedure◆ a control experiment◆ standardisation of any solution where the accuracy of the concentration is crucial in an analysis	1

Section	Expected response	Mark allocation
Data collection and handling	Experimental results providing evidence that the procedure(s) has (have) been carried out in duplicate	1
Data collection and handling	A description of the correct use of the appropriate apparatus, chemicals and any other substances to achieve the required levels of precision and/or accuracy	1
Data collection and handling	All relevant raw data is recorded	1
Data collection and handling	Numerical data is appropriately presented	1
Data collection and handling	Citations and references for three sources of internet/literature data, using relevant referencing system	1
Data analysis	Analysis of data of a level of demand commensurate with Advanced Higher Chemistry. Analysis to include, as appropriate: <ul style="list-style-type: none"> ◆ values calculated correctly using chemical relationships ◆ scatter graphs or line graphs ◆ chromatograms ◆ spectra 	4
Data analysis	Calculated values stated to an appropriate number of significant figures	1
Conclusion	A valid conclusion that relates to the aim and is supported by all the data in the report	1
Analysis	A valid comparison of the experimental data with data from the internet/literature source(s), or a comparison of duplicate/control experimental data if no internet/literature source is available	1
Evaluation	Evaluation of the investigation including, as appropriate: <ul style="list-style-type: none"> ◆ evaluation of data from the internet/literature ◆ evaluative statements supported by justification ◆ quantitative treatment of uncertainties 	4
Structure	A clear and concise report with an informative title, contents page and page numbers	1
Total marks		25

The project report is submitted to Qualifications Scotland for external marking.

All marking is quality assured by Qualifications Scotland.

Assessment conditions

Time

Candidates should start their project at an appropriate point in the course. Qualifications Scotland does not prescribe a maximum time allocation for the project but it is expected that candidates will spend 10 to 15 hours on experimental work. Candidates may choose to spend additional time on experimental work.

Supervision, control and authentication

The project is conducted under some supervision and control. This means that candidates may complete part of the work outwith the learning and teaching setting.

Teachers and lecturers must make sure candidates understand the requirements of the project from the outset.

Teachers and lecturers must ensure that the project is the work of the individual candidate, for example by:

- ◆ having regular progress meetings with candidates
- ◆ conducting spot-check interviews with candidates
- ◆ regularly reviewing candidates' lab books
- ◆ completing checklists to record candidates' progress

Teachers and lecturers must exercise their professional responsibility to ensure that the project report submitted by a candidate is the candidate's own work.

Resources

There are no restrictions on the resources to which candidates may have access.

Reasonable assistance

The term 'reasonable assistance' is used to try to balance the need for support with the need to avoid giving too much assistance, for example, drawing out or teasing out points without leading candidates. Candidates sometimes get stuck at a particular part of a task. In such cases, a teacher or lecturer could assist by raising other questions that make the candidate think about the original problem, therefore giving them the opportunity to answer their own questions without supplying the actual answers.

Teachers and lecturers must be careful that the integrity of the assessment is not compromised. Centres must not provide model answers.

Evidence to be gathered

The following candidate evidence is required for this assessment:

- ◆ a project report

The project report is submitted to Qualifications Scotland, within a given timeframe, for marking.

The same project report cannot be submitted for more than one subject.

Volume

The project report should be between 2500 and 4500 words in length, excluding the title page, contents page, tables of data, graphs, diagrams, calculations, references and acknowledgements.

Candidates must include their word count on the project report flyleaf.

If the word count exceeds the maximum by more than 10%, a penalty is applied.

Grading

Candidates' overall grades are determined by their performance across the course assessment. The course assessment is graded A–D on the basis of the total mark for both course assessment components.

Grade description for C

For the award of grade C, candidates will typically have demonstrated successful performance in relation to the skills, knowledge and understanding for the course by:

- ◆ retaining knowledge and scientific skills over an extended period of time
- ◆ integrating knowledge and understanding and scientific skills acquired throughout the course
- ◆ applying knowledge and understanding and scientific skills in a variety of contexts
- ◆ applying knowledge and understanding and scientific skills to solve problems
- ◆ selecting, analysing and presenting relevant information collected through experimental, observational or research work
- ◆ reporting in a scientific manner that communicates the chemistry

Grade description for A

For the award of grade A, candidates will typically have demonstrated a consistently high level of performance in relation to the skills, knowledge and understanding for the course by:

- ◆ retaining an extensive range of knowledge and scientific skills over an extended period of time
- ◆ integrating an extensive range of knowledge and understanding and scientific skills acquired throughout the course
- ◆ applying knowledge and understanding and scientific skills in a variety of complex contexts
- ◆ integrating knowledge and understanding and scientific skills to solve problems in a variety of complex contexts
- ◆ showing proficiency in selecting, analysing and presenting relevant information, collected through experimental, observational or research work
- ◆ showing proficiency in reporting in a scientific manner that communicates the chemistry by analysing and interpreting information in a critical and scientific manner, and demonstrating depth of knowledge and understanding

Equality and inclusion

This course is designed to be as fair and as accessible as possible.

Qualifications Scotland's [equality of access to qualifications policy](#) outlines our commitment to:

- ◆ the Equality Act 2010
- ◆ anti-racist education
- ◆ protecting children's rights
- ◆ corporate parenting

Teachers and lecturers must consider the needs of individual candidates when planning learning experiences, designing learning resources, or choosing texts.

Guidance on assessment arrangements for disabled candidates and those with additional support needs is available on the [assessment arrangements web page](#).

Further information

- ◆ [Assessment arrangements](#)
- ◆ [Building the Curriculum 3–5](#)
- ◆ [Guidance on conditions of assessment for coursework](#)
- ◆ [Guide to Assessment](#)
- ◆ [Recent Qualifications Scotland research](#)
- ◆ [Remote assessment](#)
- ◆ [SCQF Handbook](#)
- ◆ [Know Your SCQF Level — Scottish Credit and Qualifications Framework](#)
- ◆ [Skills for Learning, Skills for Life and Skills for Work framework](#)

Appendix: course support notes

Introduction

These support notes are not mandatory. They provide advice and guidance to teachers and lecturers on approaches to delivering the course. Please read these course support notes in conjunction with the course specification, the specimen question paper and the project assessment task.

Approaches to learning and teaching

This section provides you with advice and guidance on learning and teaching. You should use a variety of learning and teaching approaches to allow candidates with different needs and prior attainment to demonstrate achievement. You have considerable flexibility to select contexts that stimulate and challenge candidates, offering both breadth and depth.

Discussion and questioning are effective ways of developing candidates' knowledge and understanding of chemical concepts. Teachers, lecturers and candidates should make full use of models to develop the understanding of concepts in chemistry, and use information communication technology to support learning and to process data. As well as using computers as a learning tool, computer animations and simulations can help candidates understand chemical concepts. Computer-interfacing equipment can detect changes in variables, allowing experimental results to be recorded and processed. Results can also be displayed in real time, helping to improve understanding.

Advanced Higher courses encourage independent study. Candidates should be given opportunities to work independently, collaboratively, co-operatively and as a whole class.

You should adopt a holistic approach to encourage the simultaneous development of candidates' conceptual understanding and skills. Practical and investigative skills are strongly recommended to be progressively developed throughout the course. This allows practical techniques to be introduced and practised within real-life contexts, and for candidates to make the connections between theory and practical applications. You should encourage candidates to see risk assessment as part of the planning process for any practical activity. Throughout the course, candidates should have the opportunity to assess risks and make informed decisions regarding the use of appropriate control measures. During the project, candidates must identify safety measures taken to minimise risk during their experimental work.

Although the mandatory knowledge and skills may be similar in Higher and Advanced Higher courses, there are differences in the:

- ◆ depth of underpinning knowledge and understanding
- ◆ complexity and sophistication of the applied skills
- ◆ ways that candidates learn — they take more responsibility for their learning at Advanced Higher, and work more autonomously

The mandatory content is described in four areas: inorganic chemistry; physical chemistry; organic chemistry and instrumental analysis; and researching chemistry. Centres can deliver the course content in whichever order best meets the needs of their candidates.

Partnership working can enhance the learning experience. You could invite guest speakers from, industry, further education and higher education to share their knowledge of particular aspects of chemistry.

Assessment should be integral to and improve learning and teaching. The approach should involve candidates and provide supportive feedback. Self- and peer-assessment techniques are encouraged, wherever appropriate. Assessment information should be used to set learning targets and next steps and provide supportive feedback.

Suggested activities

Examples of possible learning and teaching activities can be found in the following table.

The first column matches the 'Skills, knowledge and understanding for the course assessment' section in the course specification. The second column offers suggestions for activities that you could use to enhance teaching and learning.

All resources named were correct at the time of publication and may be subject to change.

The Strathclyde resource, and the RSC Education site, require you to register, but registration is free.

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(a) Electromagnetic radiation and atomic spectra</p> <p>Electromagnetic radiation can be described in terms of waves and characterised in terms of wavelength and/or frequency.</p> <p>The relationship between these quantities is given by $c = f\lambda$.</p> <p>The different types of radiation arranged in order of wavelength is known as the electromagnetic spectrum.</p> <p>Wavelengths of visible light are normally expressed in nanometres (nm).</p> <p>Electromagnetic radiation can be described as a wave (has a wavelength and frequency), and as a particle, and is said to have a dual nature.</p> <p>When electromagnetic radiation is absorbed or emitted by matter it behaves like a stream of particles. These particles are known as photons.</p>	<p>Frequency is often quoted in Hz, which is the same as s^{-1}.</p> <p>A variety of different resources on this topic are available from RSC Education, including:</p> <ul style="list-style-type: none">◆ a printable handout containing a chart of the electromagnetic spectrum◆ a vignette showing the quantisation of energy levels within an atom <p>Video tutorials on light and the electromagnetic spectrum are available from Khan Academy.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>A photon carries quantised energy proportional to the frequency of radiation.</p> <p>When a photon is absorbed or emitted, energy is gained or lost by electrons within the substance.</p> <p>The photons in high frequency radiation can transfer greater amounts of energy than photons in low frequency radiation.</p> <p>The energy associated with a single photon is given by:</p> $E = hf \text{ or } E = \frac{hc}{\lambda}$ <p>The energy associated with one mole of photons is given by:</p> $E = Lhf \text{ or } E = \frac{Lhc}{\lambda}$ <p>Energy is often in units of kJ mol^{-1}.</p> <p>When energy is transferred to atoms, electrons within the atoms may be promoted to higher energy levels.</p> <p>An atom emits a photon of light energy when an excited electron moves from a higher energy level to a lower energy level.</p> <p>The light energy emitted by an atom produces a spectrum that is made up of a series of lines at discrete (quantised) energy levels. This provides direct evidence for the existence of these energy levels.</p>	<p>The values for the constants are $h = 6.63 \times 10^{-34} \text{ J s}$ and $L = 6.02 \times 10^{23} \text{ mol}^{-1}$ and are given in the data booklet.</p> <p>L is Avogadro's constant and is the number of formula units in one mole of the substance. (Formula units can be atoms, molecules or groups of ions, depending on the type of bonding present.)</p> <p>To calculate the energy of one mole of photons in kJ mol^{-1}, it may be more convenient to use:</p> $E = \frac{Lhf}{1000} \text{ or } E = \frac{Lhc}{1000\lambda}$ <p>Chemguide explains the atomic emission spectrum of hydrogen.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Each element in a sample produces characteristic absorption and emission spectra. These spectra can be used to identify and quantify the element.</p>	<p>A number of resources detail activities relating to atomic emission spectroscopy, including:</p> <ul style="list-style-type: none"> ◆ RSC Education: <ul style="list-style-type: none"> — spray bottle flame test demonstration — flame test class experiment ◆ SSERC: <ul style="list-style-type: none"> — flame test demonstration — instructions for making a spectroscope from a DVD or using a smart phone <p>Khan Academy provides a video tutorial explaining the difference between atomic absorption and emission spectroscopy.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>In absorption spectroscopy, electromagnetic radiation is directed at an atomised sample. Radiation is absorbed as electrons are promoted to higher energy levels.</p> <p>An absorption spectrum is produced by measuring how the intensity of absorbed light varies with wavelength.</p> <p>In emission spectroscopy, high temperatures are used to excite the electrons within atoms.</p> <p>As the electrons drop to lower energy levels, photons are emitted.</p> <p>An emission spectrum of a sample is produced by measuring the intensity of light emitted at different wavelengths.</p> <p>In atomic spectroscopy, the concentration of an element within a sample is related to the intensity of light emitted or absorbed.</p>	<p>A number of activities demonstrate atomic absorption spectroscopy, including:</p> <ul style="list-style-type: none"> ◆ SSERC — an activity with filter paper soaked in brine used to observe the sodium absorption spectrum ◆ vapour discharge lamps or fluorescent tube lamps used to observe the emission spectrum of mercury (a series of purple lines) when viewed through a spectroscope <p>RSC Education has an interesting anecdote describing a forensic use of atomic absorption spectroscopy in an investigation of lead in hair treated with products to reduce greyness. Emission spectra can be viewed by clicking on an element on a periodic table.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(b) Atomic orbitals, electronic configurations and the periodic table</p> <p>The discrete lines observed in atomic spectra can be explained if electrons, like photons, also display the properties of both particles and waves.</p> <p>Electrons behave as standing (stationary) waves in an atom. These are waves that vibrate in time but do not move in space. There are different sizes and shapes of standing wave possible around the nucleus, known as orbitals. Orbitals can hold a maximum of two electrons.</p> <p>The different shapes of orbitals are identified as s, p, d and f (knowledge of the shape of f orbitals is not required).</p> <p>Electrons within atoms have fixed amounts of energy called quanta.</p> <p>It is possible to describe any electron in an atom using four quantum numbers:</p> <ul style="list-style-type: none"> ◆ the principal quantum number n indicates the main energy level for an electron and is related to the size of the orbital ◆ the angular momentum quantum number l determines the shape of the subshell and can have values from zero to $n - 1$ ◆ the magnetic quantum number m_l determines the orientation of the orbital and can have values between $-l$ and $+l$ ◆ the spin magnetic quantum number m_s determines the direction of spin and can have values of $+\frac{1}{2}$ or $-\frac{1}{2}$ 	<p>Animated videos explaining the behaviour of electrons as waves and particles are available online.</p> <p>A number of resources provide details of activities relating to atomic orbitals and the quantum mechanical model of the atom:</p> <ul style="list-style-type: none"> ◆ ChemTube3D, available through RSC Education, has an illustration of the shapes of atomic orbitals ◆ Chemguide has information about atomic orbitals, including electronic configurations ◆ Khan Academy has videos and tutorials giving information on: <ul style="list-style-type: none"> — the quantum mechanical model of the atom — electronic configurations — quantum numbers ◆ RSC Education has an interactive gridlocks game on subshells

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Electrons within atoms are arranged according to:</p> <ul style="list-style-type: none"> ◆ the aufbau principle — electrons fill orbitals in order of increasing energy ('aufbau' means 'building up' in German) ◆ Hund's rule — when degenerate orbitals are available, electrons fill each singly, keeping their spins parallel before spin pairing starts ◆ the Pauli exclusion principle — no two electrons in one atom can have the same set of four quantum numbers, therefore, no orbital can hold more than two electrons and these two electrons must have opposite spins <p>In an isolated atom the orbitals within each subshell are degenerate.</p> <p>The relative energies corresponding to each orbital can be represented diagrammatically using orbital box notation for the first four shells of a multi-electron atom.</p> <p>Electronic configurations using spectroscopic notation and orbital box notation can be written for elements of atomic numbers 1 to 36.</p> <p>The periodic table is subdivided into four blocks (s, p, d and f) corresponding to the outer electronic configurations of the elements within these blocks.</p>	<p>An interesting audio discussion of Pauli's exclusion principle and the life of Wolfgang Pauli is available for download from BBC Radio 4 — In Our Time.</p> <p>There are a number of online resources providing tutorial notes covering electronic configuration, spectroscopic notation and orbital box notation, including Chemistry Libretexts and Chemguide.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The variation in first, second and subsequent ionisation energies with increasing atomic number for the first 36 elements can be explained in terms of the relative stability of different subshell electronic configurations. This provides evidence for these electronic configurations. Anomalies in the trends of ionisation energies can be explained by considering the electronic configurations.</p> <p>There is a special stability associated with half-filled and full subshells. The more stable the electronic configuration, the higher the ionisation energy.</p> <p>VSEPR (valence shell electron pair repulsion) theory can be used to predict the shapes of molecules and polyatomic ions.</p> <p>The number of electron pairs surrounding a central atom can be found by:</p> <ul style="list-style-type: none"> ◆ taking the total number of valence (outer) electrons on the central atom and adding one for each atom attached ◆ adding an electron for every negative charge ◆ removing an electron for every positive charge ◆ dividing the total number of electrons by two to give the number of electron pairs 	<p>A graph of first ionisation energies against atomic number shows anomalies, which provides good evidence of s and p orbitals being filled. Chemguide explains these anomalies.</p> <p>Although valence shell electron pair repulsion (VSEPR) theory does not describe the actual molecular orbitals in a molecule, the shapes predicted are usually quite accurate. Chemguide has instructions for working out electron pairs with worked examples.</p> <p>RSC Education has a number of vignettes on bonding theory and VSEPR.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Electron pairs are negatively charged and repel each other. They are arranged to minimise repulsion and maximise separation.</p> <p>The arrangement of electron pairs around a central atom is:</p> <ul style="list-style-type: none"> ◆ linear for two electron pairs ◆ trigonal planar for three electron pairs ◆ tetrahedral for four electron pairs ◆ trigonal bipyramidal for five electron pairs ◆ octahedral for six electron pairs <p>Shapes of molecules or polyatomic ions are determined by the shapes adopted by the atoms present based on the arrangement of electron pairs. Electron dot diagrams can be used to show these arrangements.</p> <p>Electron pair repulsions decrease in strength in the order:</p> <p>non-bonding pair/non-bonding pair > non-bonding pair/bonding pair > bonding pair/bonding pair</p>	

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(c) Transition metals</p> <p>The d-block transition metals are metals with an incomplete d subshell in at least one of their ions.</p> <p>The filling of the d orbitals follows the aufbau principle, with the exception of chromium and copper atoms.</p> <p>These exceptions are due to the special stability associated with the d subshell being half-filled or completely filled.</p> <p>When atoms from the first row of the transition elements form ions, it is the 4s electrons that are lost first rather than the 3d electrons.</p> <p>An element is said to be in a particular oxidation state when it has a specific oxidation number.</p> <p>The oxidation number can be determined by the following:</p> <ul style="list-style-type: none"> ◆ uncombined elements have an oxidation number of 0 ◆ ions containing single atoms have an oxidation number that is the same as the charge on the ion ◆ in most of its compounds, oxygen has an oxidation number of -2 ◆ in most of its compounds, hydrogen has an oxidation number of +1 ◆ the sum of all the oxidation numbers of all the atoms in a neutral compound must add up to zero ◆ the sum of all the oxidation numbers of all the atoms in a polyatomic ion must be equal to the charge on the ion 	<p>A display of sample bottles containing salts or compounds of the first 30 elements arranged on a large (A1 or A2 size) periodic table show that only the d-block compounds are coloured. Candidates may also notice that zinc compounds are white, indicating that, although lying in the central region of the periodic table, zinc is different from the transition metals. Scandium is also different since it forms only the 3⁺ ion, which has no d electrons.</p> <p>Khan Academy has a number of video tutorials on transition metals and electronic configurations of d-block elements.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>A transition metal can have different oxidation states in its compounds.</p> <p>Compounds of the same transition metal in different oxidation states may have different colours.</p> <p>Oxidation can be defined as an increase in oxidation number. Reduction can be considered as a decrease in oxidation number.</p> <p>Changes in oxidation number of transition metal ions can be used to determine whether oxidation or reduction has occurred.</p> <p>Compounds containing metals in high oxidation states are often oxidising agents, whereas compounds with metals in low oxidation states are often reducing agents.</p> <p>Ligands may be negative ions or molecules with non-bonding pairs of electrons that they donate to the central metal atom or ion, forming dative covalent bonds.</p> <p>Ligands can be classified as monodentate, bidentate, up to hexadentate.</p> <p>It is possible to deduce the ligand classification from a formula or structure of the ligand or complex.</p>	<p>A number of resources provide instructions for experiments involving oxidation states of transition metals and transition metal complexes, including:</p> <ul style="list-style-type: none"> ◆ RSC Education: <ul style="list-style-type: none"> — oxidation states of vanadium — transition elements — a microscale investigation — preparation of nickel complexes (<i>Skills' block 2</i>, page 10) — complexes of cobalt (<i>Discovery block 4</i>, page 20) — microscale iron drops practical ◆ SSERC: <ul style="list-style-type: none"> — oxidation states of vanadium — oxidation states of manganese — colour change chameleon — ligands of copper complexes — copper amino complexes ◆ Science in School: <ul style="list-style-type: none"> — a redox reaction using lollipops

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The total number of bonds from the ligands to the central transition metal is known as the coordination number.</p> <p>Names and formulae can be written according to IUPAC rules for complexes containing:</p> <ul style="list-style-type: none"> ◆ central metals that obey the normal IUPAC rules ◆ copper (cuprate) and iron (ferrate) ◆ ligands, including water, ammonia, halogens, cyanide, hydroxide, and oxalate <p>In a complex of a transition metal, the d orbitals are no longer degenerate.</p> <p>Splitting of d orbitals to higher and lower energies occurs when the electrons present in approaching ligands cause the electrons in the orbitals lying along the axes to be repelled.</p> <p>Ligands that cause a large difference in energy between subsets of d orbitals are strong field ligands. Weak field ligands cause a small energy difference.</p> <p>Ligands can be placed in an order of their ability to split d orbitals. This is called the spectrochemical series.</p>	<p>RSC Education has a gridlocks game to aid revision of shapes of complex ions and coordination numbers.</p> <p>Nomenclature of Inorganic Chemistry (Red Book) has information about IUPAC rules for naming complexes.</p> <p>Candidates can investigate the spectrochemical series and discover how the position of a ligand in the series may affect the colour of the complex. RSC <i>Education in Chemistry</i> details an experiment that can be used to demonstrate the spectrochemical series.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Colours of many transition metal complexes can be explained in terms of d-d transitions.</p> <p>Light is absorbed when electrons in a lower energy d orbital are promoted to a d orbital of higher energy.</p> <p>If light of one colour is absorbed, then the complementary colour will be observed.</p> <p>Electrons transition to higher energy levels when energy corresponding to the ultraviolet or visible regions of the electromagnetic spectrum is absorbed.</p>	<p>Chemguide explains colours of transition metal complexes.</p> <p>RSC Education hosts an exciting Nuffield Foundation experiment involving transition metal ions in coloured glass as an interesting introduction, and looks at an everyday life application of the use of transition metal chemistry.</p> <p>STEM Learning details the RSC <i>Classic Chemistry Demonstrations No. 93</i>, page 261, that shows different colours of nickel complexes with water and ethylenediamine as ligands in different ratios.</p> <p>UV-visible spectrometers and colorimeters measure the intensity of radiation transmitted through a sample, and compares this with the intensity of incident radiation. Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, details an experiment to determine the manganese content in steel using the practical technique of colorimetry.</p> <p>The wavelength ranges are approximately 200–400 nm for ultraviolet and 400–700 nm for visible light.</p>

Inorganic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Transition metals and their compounds can act as catalysts.</p> <p>Heterogeneous catalysts are in a different state to the reactants.</p> <p>Heterogeneous catalysis can be explained in terms of the formation of activated complexes and the adsorption of reactive molecules onto active sites. The presence of unpaired d electrons or unfilled d orbitals is thought to allow activated complexes to form. This can provide reaction pathways with lower activation energies compared to the uncatalysed reaction.</p> <p>Homogeneous catalysts are in the same state as the reactants.</p> <p>Homogeneous catalysis can be explained in terms of changing oxidation states with the formation of intermediate complexes.</p>	<p>Chemguide has a range of information on catalysts starting with an introduction to catalysts.</p> <p>The demonstration of cobalt(II) chloride as a homogeneous transition metal catalyst in the oxidation of Rochelle salt may have been carried out at Higher, but can now be discussed in terms of oxidation states. There are a number of online resources detailing the instructions for this, including:</p> <ul style="list-style-type: none"> ◆ RSC Education: <ul style="list-style-type: none"> — a visible activated complex — catalysts in reactions ◆ SSERC: <ul style="list-style-type: none"> — catalysts at work

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(a) Chemical equilibrium</p> <p>A chemical reaction is in equilibrium when the composition of the reactants and products remains constant indefinitely.</p> <p>The equilibrium constant (K) characterises the equilibrium composition of the reaction mixture.</p> <p>For the general reaction $aA + bB \rightleftharpoons cC + dD$ the equilibrium expression is:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>$[A], [B], [C]$ and $[D]$ are the equilibrium concentrations of A, B, C and D and a, b, c and d are the stoichiometric coefficients in the balanced reaction equation.</p> <p>The value of equilibrium constants can be calculated.</p> <p>The value of an equilibrium constant indicates the position of equilibrium.</p> <p>Equilibrium constants have no units.</p>	<p>You should reinforce the links between equilibrium where applicable in the 'Organic chemistry and instrumental analysis', and 'Researching chemistry' areas.</p> <p>Chemguide's An introduction to equilibrium has a useful recap of prior knowledge of equilibrium.</p> <p>RSC Education, Advanced starters for ten: section 3 'Equilibria', has a selection of easily editable short quizzes and activities.</p> <p>A number of tutorials on equilibrium constant are available online, for example:</p> <ul style="list-style-type: none"> ◆ Khan Academy — The equilibrium constant K ◆ Chemguide — Equilibrium constants: K_c <p>Partition coefficients could be included as a specific example of an equilibrium constant. RSC Education has details of a practical on the partition of iodine across two immiscible liquids.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The concentrations of pure solids and pure liquids at equilibrium are taken as constant and given a value of 1 in the equilibrium expression.</p> <p>The numerical value of the equilibrium constant depends on the reaction temperature and is independent of concentration and/or pressure.</p> <p>For endothermic reactions, a rise in temperature causes an increase in K and the yield of the product is increased.</p> <p>For exothermic reactions, a rise in temperature causes a decrease in K and the yield of the product is decreased.</p> <p>The presence of a catalyst does not affect the value of the equilibrium constant.</p> <p>In water and aqueous solutions there is an equilibrium between the water molecules and hydronium (hydrogen) and hydroxide ions.</p> <p>This ionisation of water can be represented by:</p> $\text{H}_2\text{O}(\ell) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{OH}^-(\text{aq})$	<p>Introduce thin-layer chromatography (TLC) and solvent extraction activities to reinforce the concept of equilibria through practical applications and to reinforce links to the 'Researching chemistry' area.</p> <p>Chemguide explains the relationship between equilibrium constant and Le Chatelier's principle.</p> <p>Chemguide has a good explanation about the ionic product of water.</p> <p>RSC Education, Advanced starters for ten: section 4, 'Acids and bases', has a selection of easily editable short quizzes and activities relating to the ionic product of water.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>$\text{H}_3\text{O}^+(\text{aq})$ represents a hydronium ion, a hydrated proton. A shorthand representation of $\text{H}_3\text{O}^+(\text{aq})$ is $\text{H}^+(\text{aq})$.</p> <p>Water is amphoteric (can react as an acid and a base).</p> <p>The dissociation constant for the ionisation of water is known as the ionic product and is represented by K_w:</p> $K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$ <p>The value of the ionic product varies with temperature.</p> <p>At 25°C the value of K_w is approximately 1×10^{-14}.</p> <p>The relationship between pH and the hydrogen ion concentration is given by:</p> $\text{pH} = -\log_{10} [\text{H}_3\text{O}^+] \text{ and } [\text{H}_3\text{O}^+] = 10^{-\text{pH}}$ <p>In water and aqueous solutions with a pH value of 7 the concentrations of $\text{H}_3\text{O}^+(\text{aq})$ and $\text{OH}^-(\text{aq})$ are both $10^{-7} \text{ mol l}^{-1}$ at 25°C.</p>	<p>There are a number of virtual lab simulations on acids and bases available, including:</p> <ul style="list-style-type: none"> ◆ Chemcollective — acid-base chemistry ◆ RSC Education — pH scale simulation

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>If the concentration of $\text{H}_3\text{O}^+(\text{aq})$ or the concentration of $\text{OH}^-(\text{aq})$ is known, the concentration of the other ion can be calculated using K_w or by using $\text{pH} + \text{pOH} = 14$.</p> <p>The Brønsted-Lowry definitions of acids and bases state that an acid is a proton donor and a base is a proton acceptor.</p> <p>For every acid there is a conjugate base, formed by the loss of a proton.</p> <p>For every base there is a conjugate acid, formed by the gain of a proton.</p> <p>Strong acids and strong bases are completely dissociated into ions in aqueous solution.</p> <p>Weak acids and weak bases are only partially dissociated into ions in aqueous solution.</p> <p>Examples of strong acids include hydrochloric acid, sulfuric acid and nitric acid.</p> <p>Ethanoic acid, carbonic acid and sulfurous acid are examples of weak acids.</p>	<p>Online tutorials available include:</p> <ul style="list-style-type: none"> ◆ Khan Academy: <ul style="list-style-type: none"> — pH, pOH and the pH scale — Brønsted-Lowry acid-base theory ◆ Chemguide: <ul style="list-style-type: none"> — theories of acids and bases, which provides information on Brønsted-Lowry acids and bases, conjugate acids and bases, and amphoteric substances <p>Investigate the pH of strong and weak acids and bases and different metal and non-metal hydroxide solutions.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Solutions of metal hydroxides are strong bases.</p> <p>Ammonia and amines are examples of weak bases.</p> <p>The weakly acidic nature of solutions of carboxylic acids, sulfur dioxide and carbon dioxide can be explained by reference to equations showing the equilibria.</p> <p>The weakly alkaline nature of a solution of ammonia or amines can be explained by reference to an equation showing the equilibrium.</p> <p>Equimolar solutions of weak and strong acids (or bases) have different pH values, conductivity, and reaction rates, but the stoichiometry of reactions are the same.</p> <p>The acid dissociation constant is represented by K_a :</p> $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ <p>or by:</p> <p>$\text{p}K_a$ where $\text{p}K_a = -\log_{10} K_a$</p> <p>The approximate pH of a weak acid can be calculated using:</p> $\text{pH} = \frac{1}{2}\text{p}K_a - \frac{1}{2}\log_{10} c .$	<p>Chemguide has tutorial notes available on strong and weak acids and bases.</p> <p>RSC Education, Advanced starters for ten: section 4, 'Acids and bases', offers easily editable short quizzes and activities covering acids and bases.</p> <p>To highlight the difference between strong and weak acids, one possible activity involves calculating the pH of a 0.1 mol l⁻¹ solution of a weak acid, and confirming the result by measurement. The solution can then be diluted tenfold to show that the pH rises by 0.5 rather than by 1, as it would when diluting a strong acid such as 0.1 mol l⁻¹ HCl. This is also a good opportunity for candidates to practise the technique of accurate dilution.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>A soluble salt of a strong acid and a strong base dissolves in water to produce a neutral solution.</p> <p>A soluble salt of a weak acid and a strong base dissolves in water to produce an alkaline solution.</p> <p>A soluble salt of a strong acid and a weak base dissolves in water to produce an acidic solution.</p> <p>The name of the salt produced depends on the acid and base used.</p> <p>Using the appropriate equilibria, the changes in concentrations of H_3O^+ and OH^- ions of salt solutions can be explained.</p>	<p>RSC Education provides resources relating to the pH of salt solutions:</p> <ul style="list-style-type: none"> ◆ Titration screen experiment: Titration 2 and titration 3 provide a good explanation of pH curves ◆ Problem-based practical activities 'Acid erosion', problem 6, contains practice calculations and a practical activity relating to pH, pH curves, strong and weak acids, and $\text{p}K_a$ values. <p>Further information about pH (titration) curves, pH of salts and choice of indicator is available on Chemguide.</p> <p>ChemCollective has a concept test on acids and bases.</p> <p>Candidates can calculate the pH of a given salt solution and confirm the value obtained by measuring the pH (examples of salt solutions include: sodium carbonate; sodium sulphite; sodium stearate; ammonium chloride; ammonium nitrate). Candidates can carry out titration experiments measuring the pH after each small addition of an acid or base, and plotting the results on a graph to create pH curves. Data logging equipment can be used and the pH curves plotted can be used to introduce indicator solutions.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>A buffer solution is one in which the pH remains approximately constant when small amounts of acid, base or water are added.</p> <p>An acid buffer consists of a solution of a weak acid and one of its salts made from a strong base.</p> <p>In an acid buffer solution the weak acid provides hydrogen ions when these are removed by the addition of a small amount of base. The salt of the weak acid provides the conjugate base, which can absorb excess hydrogen ions produced by the addition of a small amount of acid.</p> <p>A basic buffer consists of a solution of a weak base and one of its salts.</p> <p>In a basic buffer solution the weak base removes excess hydrogen ions, and the conjugate acid provided by the salt supplies hydrogen ions when these are removed.</p>	<p>Candidates can prepare buffer solutions, calculate their pH and confirm the value obtained by measuring the pH. A pH meter could be first calibrated using a buffer solution, and then used to measure the pH, and so provide a practical application of the use of a buffer solution.</p> <p>Khan Academy has a range of resources including:</p> <ul style="list-style-type: none"> ◆ pH and pKa relationship in buffers ◆ buffer solution pH calculations ◆ introduction to buffer systems, which regulate pH in blood

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>An approximate pH of an acid buffer solution can be calculated from its composition and from the acid dissociation constant:</p> $\text{pH} = \text{p}K_a - \log_{10} \frac{[\text{acid}]}{[\text{salt}]}$ <p>Indicators are weak acids for which the dissociation can be represented as:</p> $\text{HIn}(\text{aq}) + \text{H}_2\text{O}(\ell) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{In}^-(\text{aq})$ <p>The acid indicator dissociation constant is represented as K_{In} and is given by the following expression:</p> $K_{\text{In}} = \frac{[\text{H}_3\text{O}^+][\text{In}^-]}{[\text{HIn}]}$ <p>In aqueous solution the colour of an acid indicator is distinctly different from that of its conjugate base.</p>	<p>Candidates can determine the pH range over which an indicator changes colour. Natural indicators extracted from plants could be used.</p> <p>Use pH curves to explain how appropriate indicators for titration reactions are selected. Khan Academy has a video explaining titration curves and acid-base indicators.</p> <p>Demonstrate the colour changes of the various indicators present in Universal Indicator with an effervescent rainbow. Instructions are available through RSC Education.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The colour of the indicator is determined by the ratio of $[\text{HIn}]$ to $[\text{In}^-]$.</p> <p>The theoretical point at which colour change occurs is when $[\text{H}_3\text{O}^+] = K_{\text{In}}$.</p> <p>The colour change is assumed to be distinguishable when $[\text{HIn}]$ and $[\text{In}^-]$ differ by a factor of 10.</p> <p>The pH range over which a colour change occurs can be estimated by the expression:</p> $\text{pH} = \text{p}K_{\text{In}} \pm 1$ <p>Suitable indicators can be selected from pH data, including titration curves.</p>	

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(b) Reaction feasibility</p> <p>The standard enthalpy of formation, ΔH°_f, is the enthalpy change when one mole of a substance is formed from its elements in their standard states.</p> <p>The standard state of a substance is its most stable state at a pressure of 1 atmosphere and at a specified temperature, usually taken as 298 K.</p> <p>The standard enthalpy of a reaction can be calculated from the standard enthalpies of formation of the reactants and products:</p> $\Delta H^\circ = \sum \Delta H^\circ_f(\text{products}) - \sum \Delta H^\circ_f(\text{reactants})$ <p>The entropy (S) of a system is a measure of the degree of disorder of the system.</p> <p>The greater the degree of disorder, the greater the entropy.</p> <p>Solids have low disorder and gases have high disorder.</p> <p>Entropy increases as temperature increases.</p> <p>There is a rapid increase in entropy at the melting point of a substance and an even more rapid and larger change in entropy at the boiling point.</p>	<p>RSC Education, Advanced starters for ten: section 11, offers editable lesson resources on thermodynamic definitions, enthalpy, entropy and free energy.</p> <p>Khan Academy has a video that introduces entropy.</p> <p>Chemguide has useful information relating to entropy:</p> <ul style="list-style-type: none"> ◆ an introduction to entropy ◆ taking entropy changes further, which links to free energy <p>An audio discussion of the second law of thermodynamics is available for download from BBC Radio 4 — In Our Time.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The second law of thermodynamics states that the total entropy of a reaction system and its surroundings always increases for a spontaneous process.</p> <p>Heat energy released by the reaction system into the surroundings increases the entropy of the surroundings.</p> <p>Heat energy absorbed by the reaction system from the surroundings decreases the entropy of the surroundings.</p> <p>The third law of thermodynamics states that the entropy of a perfect crystal at 0 K is zero.</p> <p>The standard entropy of a substance is the entropy value for the substance in its standard state.</p> <p>The change in standard entropy for a reaction system can be calculated from the standard entropies of the reactants and products:</p> $\Delta S^{\circ} = \sum S^{\circ}(\text{products}) - \sum S^{\circ}(\text{reactants})$	<p>RSC Education has details for the endothermic solid-solid reaction between barium hydroxide and ammonium chloride. Chemicool has ideas and examples of other spontaneous endothermic reactions.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The change in free energy for a reaction is related to the enthalpy and entropy changes:</p> $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ <p>If the change in free energy (ΔG°) between reactants and products is negative, a reaction may occur and the reaction is said to be feasible. A feasible reaction is one that tends towards the products rather than the reactants. This does not give any indication of the rate of the reaction.</p> <p>The standard free energy change for a reaction can be calculated from the standard free energies of formation of the reactants and products using the relationship:</p> $\Delta G^\circ = \sum \Delta G^\circ(\text{products}) - \sum \Delta G^\circ(\text{reactants})$ <p>The feasibility of a chemical reaction under standard conditions can be predicted from the calculated value of the change in standard free energy (ΔG°).</p>	<p><i>Education in Chemistry</i> magazine provides an article outlining an experiment involving exploding soap bubbles that can be used to link between entropy changes and free energy by calculating the entropy change for the reaction of methane and oxygen.</p> <p>Chemguide provides an introduction to Gibbs free energy.</p> <p>RSC Education has two problem-solving activities:</p> <ul style="list-style-type: none"> ◆ Thermodynamics ◆ What makes it go?

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The temperatures at which a reaction may be feasible can be estimated by considering the range of values of T for which $\Delta G^\circ < 0$.</p> <p>Under non-standard conditions any reaction is feasible if ΔG is negative.</p> <p>At equilibrium, $\Delta G = 0$.</p> <p>A reversible reaction will proceed spontaneously until the composition is reached where $\Delta G = 0$.</p> <p>(c) Kinetics</p> <p>The rate of a chemical reaction normally depends on the concentrations of the reactants.</p> <p>Orders of reaction are used to relate the rate of a reaction to the reacting species.</p> <p>If changing the concentration of a reactant A has no effect on the rate of the reaction, then the reaction is zero order with respect to A.</p> <p>If doubling the concentration of a reactant A doubles the rate of the reaction, then the reaction is first order with respect to A. The rate can be expressed as:</p> <p>rate = $k[A]$ where k is the rate constant and $[A]$ is the concentration of reactant A in mol l^{-1}</p>	<p>Carry out an experiment to verify a thermodynamic prediction using, for example, $\text{NaHCO}_3(\text{s})$.</p> <p>RSC Education, Advanced starters for ten: section 2, offers editable lesson resources on kinetics.</p> <p>Chemguide has information on orders of reaction and rate equations and order of reaction and organic mechanisms.</p> <p>A video tutorial is also available on Khan Academy.</p>

Physical chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>If doubling the concentration of a reactant A increases the rate of the reaction fourfold, then the reaction is second order with respect to A. The rate can be expressed as:</p> $\text{rate} = k[A]^2$ <p>The order of a reaction with respect to any one reactant is the power to which the concentration of that reactant is raised in the rate equation.</p> <p>The overall order of a reaction is the sum of the powers to which the concentrations of the reactants are raised in the rate equation.</p> <p>The order of a reaction can only be determined from experimental data.</p> <p>The rate equation and the rate constant, including units, can be determined from initial rate data for a series of reactions in which the initial concentrations of reactants are varied. These can be zero, first, second or third order.</p> <p>Reactions usually occur by a series of steps called a reaction mechanism.</p> <p>The rate of reaction is dependent on the slowest step, which is called the 'rate determining step'.</p> <p>Experimentally determined rate equations can be used to determine possible reaction mechanisms.</p>	<p>A number of instructions for practical activities are available:</p> <ul style="list-style-type: none"> ◆ RSC Education's problem-based practical activities provides problem-solving activities and experimental details for the propanone and iodine reaction. ◆ SSERC describes an experiment to determine the rate constant and order of reaction using bleach and blue food dye, and also provides an opportunity to introduce the practical technique of colorimetry. ◆ The University of Strathclyde has details of an experiment, determination of the rate of a reaction that includes detailed kinetics information, as well as the experimental procedure.

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(a) Molecular orbitals</p> <p>VSEPR cannot explain the bonding in all compounds. Molecular orbital theory can provide an explanation for more complex molecules.</p> <p>Molecular orbitals form when atomic orbitals combine. The number of molecular orbitals formed is equal to the number of atomic orbitals that combine. The combination of two atomic orbitals results in the formation of a bonding molecular orbital and an antibonding orbital. The bonding molecular orbital encompasses both nuclei. The attraction of the positively charged nuclei and the negatively charged electrons in the bonding molecular orbital is the basis of bonding between atoms. Each molecular orbital can hold a maximum of two electrons.</p> <p>In a non-polar covalent bond, the bonding molecular orbital is symmetrical about the midpoint between two atoms. Polar covalent bonds result from bonding molecular orbitals that are asymmetric about the midpoint between two atoms. The atom with the greater value for electronegativity has the greater share of the bonding electrons. Ionic compounds are an extreme case of asymmetry, with the bonding molecular orbitals being almost entirely located around just one atom, resulting in the formation of ions.</p> <p>Molecular orbitals that form by end-on overlap of atomic orbitals along the axis of the covalent bond are called sigma (σ) molecular orbitals or sigma bonds.</p>	<p>There are websites available with information and animations showing sigma bonds, pi bonds and hybridisation.</p> <p>RSC Education has a series of vignettes covering molecular orbitals, hybridisation, aromaticity and conjugation.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Molecular orbitals that form by side-on overlap of parallel atomic orbitals that lie perpendicular to the axis of the covalent bond are called pi (π) molecular orbitals or pi bonds.</p> <p>The electronic configuration of an isolated carbon atom cannot explain the number of bonds formed by carbon atoms in molecules. The bonding and shape of molecules of carbon can be explained by hybridisation.</p> <p>Hybridisation is the process of mixing atomic orbitals within an atom to generate a set of new atomic orbitals called hybrid orbitals. These hybrid orbitals are degenerate.</p> <p>In alkanes, the 2s orbital and the three 2p orbitals of carbon hybridise to form four degenerate sp^3 hybrid orbitals. These adopt a tetrahedral arrangement. The sp^3 hybrid orbitals overlap end-on with other atomic orbitals to form σ bonds.</p> <p>The bonding in alkenes can be described in terms of sp^2 hybridisation. The 2s orbital and two of the 2p orbitals hybridise to form three degenerate sp^2 hybrid orbitals. These adopt a trigonal planar arrangement. The hybrid sp^2 orbitals overlap end-on to form σ bonds. The remaining 2p orbital on each carbon atom of the double bond is unhybridised and lies perpendicular to the axis of the σ bond. The unhybridised p orbitals overlap side-on to form π bonds.</p>	<p>ChemTube3D, available through RSC Education, contains interactive 3D models for some important organic molecules including methane, ethane, ethyne and benzene. The model view can be altered to show the hybrid orbitals.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The bonding in benzene and other aromatic systems can be described in terms of sp^2 hybridisation. The six carbon atoms in benzene are arranged in a cyclic structure with σ bonds between the carbon atoms. The unhybridised p orbitals on each carbon atom overlap side-on to form a π molecular system, perpendicular to the plane of the σ bonds. This π molecular system extends across all six carbon atoms. The electrons in this system are delocalised.</p> <p>The bonding in alkynes can be described in terms of sp hybridisation. The 2s orbital and one 2p orbital of carbon hybridise to form two degenerate hybrid orbitals. These adopt a linear arrangement. The hybrid sp orbitals overlap end-on to form σ bonds. The remaining two 2p orbitals on each carbon atom lie perpendicular to each other and to the axis of the σ bond. The unhybridised p orbitals overlap side-on to form two π bonds.</p> <p>Molecular orbital theory can be used to explain why organic molecules are colourless or coloured. Electrons fill bonding molecular orbitals, leaving higher energy antibonding orbitals unfilled. The highest bonding molecular orbital containing electrons is called the highest occupied molecular orbital (HOMO). The lowest antibonding molecular orbital is called the lowest unoccupied molecular orbital (LUMO).</p> <p>Absorption of electromagnetic energy can cause electrons to be promoted from HOMO to LUMO.</p> <p>Most organic molecules appear colourless because the energy difference between HOMO and LUMO is relatively large. This results in absorption of light from the ultraviolet region of the spectrum.</p>	<p>Khan Academy has a series of videos covering molecular orbitals, HOMO and LUMO, and UV/Vis spectroscopy in organic molecules as well as explaining the link between conjugation and colour in organic molecules. These videos compare the absorptions of molecules with different degrees of conjugated systems.</p> <p>ChemTube3D, available through RSC Education, has many useful resources including a graphic showing the difference in energy between HOMO and LUMO in linear polyenes as well as 3D models showing the conjugation in a number of dyes.</p> <p>A PowerPoint presentation introducing molecules and colour is a resource produced by Education Scotland, and available on the Science NQ GLOW portal.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Some organic molecules contain chromophores. A chromophore is a group of atoms within a molecule that is responsible for absorption of light in the visible region of the spectrum. Light can be absorbed when electrons in a chromophore are promoted from the HOMO to the LUMO.</p> <p>Chromophores exist in molecules containing a conjugated system — a system of adjacent unhybridised p orbitals that overlap side-on to form a molecular orbital across a number of carbon atoms. Electrons within this conjugated system are delocalised. Molecules with alternating single and double bonds, and aromatic molecules have conjugated systems.</p> <p>The more atoms in the conjugated system the smaller the energy gap between HOMO and LUMO. A lower frequency of light (longer wavelength, lower energy) is absorbed by the compound. When the wavelength of light absorbed is in the visible region, the compound will exhibit the complementary colour.</p>	<p>Chemistry in your cupboard is a resource available on RSC Education, and links the action of the stain remover Vanish to the structure of coloured organic molecules, providing a real-life example of the importance of understanding the chemistry of colour.</p> <p>Colourful Chemistry infographics from RSC Education has a visually stimulating and informative infographic about colours of organic molecules including the colours of autumn leaves, poinsettia plants and glow sticks.</p> <p>The Science of Sunscreen infographic from RSC Education has information about organic molecules with conjugated systems that are used to absorb UV light.</p> <p>Compounds highlighting the effect of increasing the length of the conjugated system on the colour can be compared. For example, vitamin A (yellow) can be compared with β-carotene (orange).</p> <p>Ninhydrin reacts with amino acids and forms a highly conjugated product that absorbs light in the visible region, and an intense purple colour (λ_{max} 750 nm) is observed. This is used in the detection of amino acids both as a locating agent in chromatography and in detecting latent fingerprints in crime scenes. Molecule of the Month: April 2018 outlines the steps involved in formation of one of these conjugated products.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
	<p>Explain the colours observed for the indicator phenolphthalein by looking at the degree of conjugation in the molecule in both acid and in base conditions. Structures of the two forms of phenolphthalein can be seen on the Imperial College website.</p> <p>An experiment to synthesise phenolphthalein and its derivatives is available through the University of Strathclyde website.</p> <p>Prepare a variety of dyes and examine their structures to identify the chromophore. RSC Education has instructions for the microscale synthesis of an azo dye and the microscale synthesis of indigo-dye. The University of Strathclyde has instructions for the synthesis of methyl orange.</p> <p>A number of online resources allow complementary colours to be explained by demonstrating colour mixing.</p> <p>Candidates can use simple spectrometers made from DVDs or using a smart phone — available from SSERC — to view light transmitted or reflected by coloured compounds. SSERC has instructions for making spectrometers.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(b) Synthesis</p> <p>When an organic reaction takes place, bonds in the reactant molecules are broken and bonds in the product molecules are made. The process of bond breaking is known as bond fission.</p> <p>There are two types of bond fission, homolytic and heterolytic.</p> <p>Homolytic fission:</p> <ul style="list-style-type: none"> ◆ results in the formation of two neutral radicals ◆ occurs when each atom retains one electron from the σ covalent bond and the bond breaks evenly ◆ normally occurs when non-polar covalent bonds are broken <p>Reactions involving homolytic fission tend to result in the formation of very complex mixtures of products, making them unsuitable for organic synthesis.</p> <p>Heterolytic fission:</p> <ul style="list-style-type: none"> ◆ results in the formation of two oppositely charged ions ◆ occurs when one atom retains both electrons from the σ covalent bond and the bond breaks unevenly ◆ normally occurs when polar covalent bonds are broken <p>Reactions involving heterolytic fission tend to result in far fewer products than reactions involving homolytic fission, and so are better suited for organic synthesis.</p>	<p>Sacramento City College offers a lesson that covers homolytic and heterolytic fission, along with curly arrow notation, and definitions of nucleophiles and electrophiles.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The movement of electrons during bond fission and bond making can be represented using curly arrow notation where:</p> <ul style="list-style-type: none"> ◆ a single-headed arrow indicates the movement of a single electron ◆ a double-headed arrow indicates the movement of an electron pair ◆ the tail of the arrow shows the source of the electron(s) ◆ the head of the arrow indicates the destination of the electron(s) ◆ two single-headed arrows starting at the middle of a covalent bond indicate homolytic bond fission is occurring ◆ a double-headed arrow starting at the middle of a covalent bond indicates heterolytic bond fission is occurring ◆ an arrow drawn with the head pointing to the space between two atoms indicates that a covalent bond will be formed between those two atoms <p>In reactions involving heterolytic bond fission, attacking groups are classified as nucleophiles or electrophiles.</p>	<p>Chemguide has information on use of curly arrows.</p> <p>RSC Mechanism Inspector has information and interactive activities relating to single and double-headed curly arrows, electrophiles and nucleophiles.</p> <p><i>Education in Chemistry</i> magazine provides an article, 'End curly arrow anxiety', which has a downloadable exercise to practice curly arrow mechanisms.</p> <p>RSC Education has a set of activity sheets to aid understanding of curly arrows and reaction mechanisms.</p> <p>The University of Southampton has created a set of exam-style self-assessment questions covering various aspects of organic reactions and mechanisms.</p> <p>RSC Mechanism Inspector has information and interactive activities relating to electrophiles and nucleophiles.</p> <p>RSC Education has a series of vignettes covering mechanism theory including electrophiles, nucleophiles and curly arrow notation.</p>


Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Nucleophiles are:</p> <ul style="list-style-type: none"> ◆ negatively charged ions or neutral molecules that are electron rich, such as Cl^-, Br^-, OH^-, CN^-, NH_3 and H_2O ◆ attracted towards atoms bearing a partial (δ^+) or full positive charge ◆ capable of donating an electron pair to form a new covalent bond <p>Electrophiles are:</p> <ul style="list-style-type: none"> ◆ positively charged ions or neutral molecules that are electron deficient, such as H^+, NO_2^+ and SO_3 ◆ attracted towards atoms bearing a partial (δ^-) or full negative charge ◆ capable of accepting an electron pair to form a new covalent bond <p>The following reaction types can be identified from a chemical equation:</p> <ul style="list-style-type: none"> ◆ substitution ◆ addition ◆ elimination ◆ condensation ◆ hydrolysis ◆ oxidation ◆ reduction ◆ neutralisation 	<p>RSC Education, Starters for ten: chapters 1–12, section 6 'Organic Chemistry', offers a selection of easily editable short quizzes and activities covering curly arrows and electrophiles and nucleophiles.</p> <p>National 5 and Higher courses cover some of the reaction types. Candidates should revise these in the context of this course.</p> <p>Organic Chemistry infographics from RSC Education has a visually stimulating and informative infographic relating different reaction types in organic chemistry.</p> <p>It is important that you give candidates many varied, real-life contexts for these reactions. Similarities between the different reaction types should be reinforced and opportunities given to make connections between reaction types and to develop skills that enable synthetic routes to be devised for given products.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Synthetic routes can be devised, with no more than three steps, from a given reactant to a final product.</p> <p>The possible reactions of a particular molecule can be deduced by looking at the structural formula.</p> <p>The structure of any molecule can be drawn as a full, shortened or skeletal structural formula.</p> <p>In a skeletal structural formula, neither the carbon atoms, nor any hydrogens attached to the carbon atoms, are shown. The presence of a carbon atom is implied by a 'kink' in the carbon backbone, and at the end of a line.</p> <p>Given a full or shortened structural formula for a compound, the skeletal structural formula can be drawn.</p> <p>Given a skeletal structural formula for a compound, the full or shortened structural formula can be drawn.</p> <p>Molecular formulae can be written from a full, shortened or skeletal structural formula.</p>	<p>Which Compound? Which Route? is available on RSC Education and has an activity to plan a synthetic route for a drug.</p> <p>RSC Education has details on Synthesis Explorer that allow synthetic routes to be devised, and shows details of reagents involved.</p> <p>Chemguide provides information about skeletal formula.</p> <p>Organic chemistry infographics from RSC Education has a visually stimulating and informative infographic showing the different types of organic formulae.</p> <p>Molecular drawing packages such as ChemSketch (a free downloadable application) can be set to display structures in skeletal representation if required. Wireframe, stick, ball and stick, and space-filling representations should all be familiar. Molecules can be rotated around the <i>x</i>, <i>y</i> and <i>z</i> axes to align any chosen bond horizontally or vertically; to align any three atoms in a given plane; to zoom in and out; and to switch on and off atom labels. Molecules sketched in 2D mode can be converted into 3D representations in ChemSketch.</p> <p>Candidates can create and manipulate 3D representations of relatively small molecules (fewer than 10 carbon atoms) containing common functional groups using Molymods or other molecular model kits.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Straight and branched chain alkanes; alkenes; alcohols; carboxylic acids; aldehydes and ketones; haloalkanes; and ethers can be systematically named, indicating the position of the functional group where appropriate, from structural formulae containing no more than eight carbon atoms in their longest chain. Straight chain esters can be systematically named from the names of their parent alcohol and carboxylic acid or their structural formula.</p> <p>Molecular formulae can be written and structural formulae drawn from systematic names of straight and branched chain alkanes; alkenes; alcohols; carboxylic acids; aldehydes and ketones; haloalkanes; and ethers containing no more than eight carbon atoms in their longest chain. Molecular formulae can be written and structural formulae drawn for esters from the systematic name or the structural formulae of their parent alcohol and carboxylic acid.</p> <p>Haloalkanes (alkyl halides) are substituted alkanes in which one or more of the hydrogen atoms is replaced with a halogen atom.</p>	<p>ChemSpider has a searchable library of molecules that can be viewed as skeletal formula or as 3D molecules.</p> <p>RSC Education's Gridlocks has some activities that may help to revise naming rules from previous courses.</p> <p>Chemguide has useful information that explains the naming of all types of organic compound.</p> <p>Orgchem101, produced by the University of Ottawa, has an interactive organic nomenclature quiz.</p> <p>A podcast on the environmental significance and the chemistry of haloalkanes is available from RSC Education.</p> <p>RSC Education details a Nuffield Foundation experiment to synthesise bromoethane from ethanol in a test tube.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Monohaloalkanes:</p> <ul style="list-style-type: none"> ◆ contain only one halogen atom ◆ can be classified as primary, secondary or tertiary according to the number of alkyl groups attached to the carbon atom containing the halogen atom ◆ take part in elimination reactions to form alkenes using a strong base, such as potassium or sodium hydroxide in ethanol ◆ take part in nucleophilic substitution reactions with: <ul style="list-style-type: none"> — aqueous alkalis to form alcohols — alcoholic alkoxides to form ethers — ethanolic cyanide to form nitriles (chain length increased by one carbon atom) that can be hydrolysed to carboxylic acids <p>A monohaloalkane can take part in nucleophilic substitution reactions by one of two different mechanisms.</p> <p>S_N1 is a nucleophilic substitution reaction with one species in the rate determining step and occurs in a minimum of two steps via a trigonal planar carbocation intermediate.</p>	<p>RSC Education, Starters for ten: chapters 1–12, section 6 'Organic Chemistry', offers a selection of easily editable short quizzes and activities relating to haloalkanes.</p> <p>Chemguide has a description of the different kinds of haloalkanes as well as elimination and nucleophilic reactions.</p> <p><i>Education in Chemistry</i> magazine provides an article outlining an investigation into the mechanism of the hydrolysis of 2-bromo-2-methylpropane.</p> <p>RSC Mechanism Inspector has information and interactive activities about nucleophilic substitution reactions.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>S_N2 is a nucleophilic substitution reaction with two species in the rate determining step and occurs in a single step via a single five-centred, trigonal bipyramidal transition state.</p> <p>The reaction mechanisms for S_N1 and S_N2 reactions can be represented using curly arrows. Steric hindrance and the inductive stabilisation of the carbocation intermediate can be used to explain which mechanism will be preferred for a given haloalkane.</p> <p>Alcohols are substituted alkanes in which one or more of the hydrogen atoms is replaced with a hydroxyl functional group, $-OH$ group.</p>	<p>The University of Oxford has an interactive quiz to test knowledge of nucleophilic substitution reactions and curly arrow mechanisms.</p> <p>ChemTube3D, available through RSC Education, has simple animated mechanisms and 3D models showing nucleophilic substitution reactions as well as showing more complex examples.</p> <p>Khan Academy has videos showing the mechanism of both S_N1 and S_N2 reactions.</p> <p>RSC Education, Starters for ten: chapters 1–12, section 6 'Organic Chemistry', offers a selection of easily editable short quizzes and activities relating to substitution reactions and elimination reactions of haloalkanes.</p> <p>RSC Education, Starters for ten: chapters 1–12, section 6 'Organic Chemistry', offers a selection of easily editable short quizzes and activities testing knowledge of alcohols and their reactions.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Alcohols can be prepared from:</p> <ul style="list-style-type: none"> ◆ haloalkanes by substitution ◆ alkenes by acid-catalysed hydration (addition) ◆ aldehydes and ketones by reduction using a reducing agent such as lithium aluminium hydride <p>Reactions of alcohols include:</p> <ul style="list-style-type: none"> ◆ dehydration to form alkenes using aluminium oxide, concentrated sulfuric acid or concentrated phosphoric acid ◆ oxidation of primary alcohols to form aldehydes and then carboxylic acids and secondary alcohols to form ketones, using acidified permanganate, acidified dichromate or hot copper(II) oxide ◆ formation of alcoholic alkoxides by reaction with some reactive metals such as potassium or sodium, which can then be reacted with monohaloalkanes to form ethers ◆ formation of esters by reaction with carboxylic acids using concentrated sulfuric acid or concentrated phosphoric acid as a catalyst ◆ formation of esters by reaction with acid chlorides (see diagram below) — this gives a faster reaction than reaction with carboxylic acids, and no catalyst is needed 	<p>RSC Education provides a range of activities on reactions of alcohols:</p> <ul style="list-style-type: none"> ◆ dehydration of ethanol to ethene using porcelain chips as a catalyst ◆ preparation of cyclohexene from cyclohexanol (<i>Skills – Block 1</i>, page 39) with purification by distillation and solvent extraction and testing the product for unsaturation ◆ oxidation of ethanol — the alcohol is oxidised to ethanal and, with further oxidation, to ethanoic acid ◆ the 'breathalyser' reaction is a quick demonstration of the reaction used in early forms of breathalysers ◆ a microscale oxidation of alcohols allows the difference in the oxidation reactions of primary, secondary and tertiary alcohols to be observed by the addition of acidified dichromate(VI) ◆ properties of alcohols details the reaction of sodium with ethanol ◆ Alcohols (16-19) is a game and resource based on naming, classifying and identifying the products of oxidation ◆ making esters from alcohols and acids ◆ microscale synthesis of ethyl benzoate ◆ Chemguide, which has an explanation of the properties of alcohols in relation to hydrogen bonding

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Hydroxyl groups make alcohols polar, which gives rise to hydrogen bonding. Hydrogen bonding can be used to explain the properties of alcohols including boiling points, melting points, viscosity and solubility or miscibility in water.</p> <p>Ethers can be regarded as substituted alkanes in which a hydrogen atom is replaced with an alkoxy functional group, –OR, and have the general structure R' – O – R'', where R' and R'' are alkyl groups.</p> <p>Ethers are named as substituted alkanes. The alkoxy group is named by adding the ending 'oxy' to the alkyl substituent, and this prefixes the name of the longest carbon chain.</p> <p>Ethers can be prepared in a nucleophilic substitution reaction by reacting a monohaloalkane with an alkoxide.</p> <p>Due to the lack of hydrogen bonding between ether molecules, they have lower boiling points than the corresponding isomeric alcohols.</p> <p>Methoxymethane and methoxyethane are soluble in water. Larger ethers are insoluble in water due to their increased molecular size.</p> <p>Ethers are commonly used as solvents since they are relatively inert chemically and will dissolve many organic compounds.</p>	<p>SSERC has details of the Ether Runway demonstration, which provides an interesting introduction to ethers, illustrating their flammability.</p> <p><i>ChemistryWorld</i> magazine has a podcast and transcript providing some of the history of diethyl ether.</p> <p>Khan Academy has a video that explains IUPAC naming of ethers as well as describing some of their properties. Another video describes the Williamson ether synthesis of an alcohol and alkyl halide using a strong base such as sodium hydride.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Alkenes can be prepared by:</p> <ul style="list-style-type: none"> ◆ dehydration of alcohols using aluminium oxide, concentrated sulfuric acid or concentrated phosphoric acid ◆ base-induced elimination of hydrogen halides from monohaloalkanes <p>Alkenes take part in electrophilic addition reactions with:</p> <ul style="list-style-type: none"> ◆ hydrogen to form alkanes in the presence of a catalyst ◆ halogens to form dihaloalkanes ◆ hydrogen halides to form monohaloalkanes ◆ water using an acid catalyst to form alcohols 	<p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, has details of the preparation of cyclohexene from cyclohexane using concentrated phosphoric acid and provides an opportunity to introduce the practical techniques of distillation and solvent extraction (only one extraction is carried out in this procedure).</p> <p>RSC Education has a number of experiments and activities relating to alkenes including:</p> <ul style="list-style-type: none"> ◆ dehydration of ethanol to ethene using porcelain chips as a catalyst ◆ preparation of cyclohexene from cyclohexanol with purification by distillation and solvent extraction ◆ Starters for ten: chapters 1–12, section 6 'Organic Chemistry' offers a selection of easily editable short quizzes and activities testing knowledge of electrophilic addition reactions of alkenes, including reaction mechanisms <p>RSC Mechanism Inspector has information and interactive activities about electrophilic addition reactions.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Markovnikov's rule states that when a hydrogen halide or water is added to an unsymmetrical alkene, the hydrogen atom becomes attached to the carbon with the most hydrogen atoms attached to it already. Markovnikov's rule can be used to predict major and minor products formed during the reaction of a hydrogen halide or water with alkenes.</p> <p>The reaction mechanisms for the addition of a hydrogen halide and the acid-catalysed addition of water can be represented using curly arrows and showing the intermediate carbocation. The inductive stabilisation of intermediate carbocations formed during these reactions can be used to explain the products formed.</p> <p>The reaction mechanism for the addition of a halogen can be represented using curly arrows and showing the cyclic ion intermediate.</p> <p>Carboxylic acids can be prepared by:</p> <ul style="list-style-type: none"> ◆ oxidising primary alcohols using acidified permanganate, acidified dichromate and hot copper(II) oxide ◆ oxidising aldehydes using acidified permanganate, acidified dichromate, Fehling's solution and Tollens' reagent ◆ hydrolysing nitriles, esters or amides 	<p>Khan Academy has a video that explains Markovnikov's rule, using curly arrows and inductive stabilisation of carbocation intermediates.</p> <p>Chemguide has mechanisms for electrophilic addition reactions of alkenes with hydrogen halides and halogens.</p> <p>The Khan Academy video hydration of an alkene shows the mechanism for the reaction of an alkene with water and a sulfuric acid catalyst.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Reactions of carboxylic acids include:</p> <ul style="list-style-type: none"> ◆ formation of salts by reactions with metals or bases ◆ condensation reactions with alcohols to form esters in the presence of concentrated sulfuric or concentrated phosphoric acid ◆ reaction with amines to form alkylammonium salts that form amides when heated ◆ reduction with lithium aluminium hydride to form primary alcohols 	<p>RSC Education provides a range of resources relating to carboxylic acids:</p> <ul style="list-style-type: none"> ◆ Advanced starters for ten: section 5, 'Carbonyl Chemistry' has easily editable short quizzes on reactions of carboxylic acids ◆ oxidation of ethanol experiment ◆ salicylic acid infographic provides information relating to the wide-ranging use of this carboxylic acid ◆ reactions of ethanoic acid compares the pH of ethanoic acid and hydrochloric acid as well as their reactions with magnesium and sodium carbonate ◆ microscale preparation of ethyl benzoate ◆ microscale synthesis of aspirin ◆ synthesis of aspirin, which also provides an opportunity to introduce important practical techniques including reflux, recrystallisation, melting point analysis, thin-layer chromatography and % yield calculations. A similar procedure can also be found in Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC ◆ The First Year Undergraduate Chemistry Laboratory Course Manual 2011–2012 provides experimental details that include use of important practical techniques (including reflux, vacuum filtration, recrystallisation, melting point analysis and % yield calculations) for preparation of the ester propyl ethanoate as well as the preparation of benzoic acid from methyl benzoate

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Amines are organic derivatives of ammonia in which one or more hydrogen atoms of ammonia has been replaced by an alkyl group.</p> <p>Amines can be classified as primary, secondary or tertiary according to the number of alkyl groups attached to the nitrogen atom.</p> <p>Amines react with acids to form salts.</p> <p>Primary and secondary amines, but not tertiary amines, display hydrogen bonding. As a result, primary and secondary amines have higher boiling points than isomeric tertiary amines.</p> <p>Primary, secondary and tertiary amine molecules can hydrogen-bond with water molecules, thus explaining the appreciable solubility of the shorter chain length amines in water.</p> <p>Amines like ammonia are weak bases and dissociate to a slight extent in aqueous solution. The nitrogen atom has a lone pair of electrons which can accept a proton from water, producing hydroxide ions.</p>	<p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, has details of an experiment to hydrolyse ethyl benzoate to prepare benzoic acid, and provides an opportunity to introduce important practical techniques including reflux, recrystallisation, melting point analysis, thin-layer chromatography and % yield calculations. This guide also contains experimental details to prepare ethyl ethanoate.</p> <p>RSC Education, Advanced starters for ten: section 7, 'Compounds with amine groups', has easily editable short quizzes on reactions and properties of amines.</p> <p>Chemguide has useful information about amines including descriptions of the different classifications and their properties and reactions.</p> <p>The University of Purdue has some information about amine-based drugs and discusses the solubility of some of the available forms.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Benzene (C₆H₆) is the simplest member of the class of aromatic hydrocarbons.</p> <p>The benzene ring has a distinctive structural formula. The stability of the benzene ring is due to the delocalisation of electrons in the conjugated system. The presence of delocalised electrons explains why the benzene ring does not take part in addition reactions.</p> <p>Bonding in benzene can be described in terms of sp² hybridisation, sigma and pi bonds, and electron delocalisation.</p> <p>A benzene ring in which one hydrogen atom has been substituted by another group is known as the phenyl group. The phenyl group has the formula –C₆H₅.</p> <p>Benzene rings can take part in electrophilic substitution reactions.</p> <p>Reactions at benzene rings include:</p> <ul style="list-style-type: none"> ◆ halogenation by reaction of a halogen using aluminium chloride or iron(III) chloride for chlorination and aluminium bromide or iron(III) bromide for bromination ◆ alkylation by reaction of a haloalkane using aluminium chloride ◆ nitration using concentrated sulfuric acid and concentrated nitric acid ◆ sulfonation using concentrated sulfuric acid 	<p><i>ChemistryWorld</i> magazine has a podcast and transcript providing some of the history of benzene.</p> <p>Many everyday consumer products have very distinctive smells as a result of the presence of key aromatic compounds. Create a display of household products containing these aromatic compounds to capture interest. Examples could include well known antiseptics and disinfectants containing trichlorophenol or 4-chloro-3,5-dimethylphenol, or permanent markers containing xylene or toluene.</p> <p>ChemGuide provides a good explanation about:</p> <ul style="list-style-type: none"> ◆ the bonding in benzene ◆ the modern representation of benzene ◆ electrophilic substitution reactions <p>RSC Education, Advanced starters for ten: section 6, 'Aromatic Chemistry', has easily editable short quizzes and activities on reactions and properties of aromatic compounds.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(c) Stereo chemistry</p> <p>Molecules that have the same molecular formula but different structural formulae are called isomers.</p> <p>Structural isomers occur when the atoms are bonded together in a different order in each isomer.</p>	<p>There are many important compounds that contain benzene rings. RSC Education provides details about:</p> <ul style="list-style-type: none"> ◆ aspirin ◆ paracetamol ◆ ibuprofen (Nurofen) <p>RSC Education has details of a micro-scale preparation of TCP, as well as the nitration of methyl benzoate.</p> <p>Electrophilic aromatic substitution reactions, available through the University of Strathclyde, provides details for carrying out some electrophilic aromatic substitution reactions and also provides an opportunity to introduce the practical techniques of recrystallisation, melting point analysis and thin-layer chromatography.</p> <p>Khan Academy has a short video that introduces structural and stereo isomerism.</p> <p>Candidates can create and manipulate 3D representations of relatively small molecules (fewer than 10 carbon atoms) using Molymods or other molecular model kits, to show the difference in structures of geometric and optical isomers.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Stereoisomers occur when the order of the bonding in the atoms is the same but the spatial arrangement of the atoms is different in each isomer. There are two types of stereoisomer, geometric and optical.</p> <p>Geometric isomers:</p> <ul style="list-style-type: none"> ◆ can occur when there is restricted rotation around a carbon-carbon double bond or a carbon-carbon single bond in a cyclic compound ◆ must have two different groups attached to each of the carbon atoms that make up the bond with restricted rotation ◆ can be labelled cis or trans according to whether the substituent groups are on the same side (cis) or on different sides (trans) of the bond with restricted rotation ◆ have differences in physical properties, such as melting point and boiling point ◆ can have differences in chemical properties <p>Optical isomers:</p> <ul style="list-style-type: none"> ◆ occur in compounds in which four different groups are arranged tetrahedrally around a central carbon atom (chiral carbon or chiral centre) ◆ are asymmetric ◆ are non-superimposable mirror images of each other ◆ can be described as enantiomers 	<p>RSC Education, Starters for ten: chapters 1–12, section 6 'Organic Chemistry', offers a selection of easily editable short quizzes and activities testing knowledge of structural and geometric isomers.</p> <p>Videos explaining the difference in physical and chemical properties of geometric isomers are available online.</p> <p>A useful article produced by the American Chemical Society discusses saturated, unsaturated and cis and trans fats, and some of the health concerns relating to trans fats.</p> <p><i>ChemistryWorld</i> magazine has a podcast and transcript about the anticancer drug cis-platin.</p> <p>An experiment to synthesise cis and trans complexes of cobalt is available through the University of Strathclyde.</p> <p>The Khan Academy video available through RSC Education introduces chirality and provides worked examples of molecules with and without chiral carbon atoms.</p> <p>RSC Education has details of two short activities to look at the differences in properties of the two optical isomers of limonene as well as those for carvone (caraway and spearmint). There is also an experiment to explore the optical rotation of sugars.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<ul style="list-style-type: none"> ◆ have identical physical properties, except for their effect on plane-polarised light ◆ have identical chemical properties, except when in a chiral environment such as that found in biological systems (only one optical isomer is usually present) ◆ rotate plane-polarised light by the same amount but in opposite directions and so are optically active ◆ when mixed in equal amounts are optically inactive because the rotational effect of the plane-polarised light cancels out — this is called a racemic mixture 	<p>Candidates can build their own polarimeters. The RSC describes a demonstration to show chirality in sugars, which also includes instructions for constructing a simple polarimeter.</p> <p>ChemTube3D, available through RSC Education, has a tutorial that discusses the differences between chiral and non-chiral molecules, including an activity in which the molecules can be rotated to show they are non-superimposable.</p> <p>The University of Bristol's, Molecule of the Month, July 2000, features thalidomide and discusses the uses of the drug, the enantiomeric forms and the associated consequences of the use of thalidomide in pregnant women. Limonene features as Molecule of the Month: March 2008.</p> <p>RSC Education features Chemistry in your Cupboard: Nurofen, which is normally sold as a mixture of two optical isomers, one of which is an effective pain-killing drug and the other of which is inactive.</p> <p>One enantiomer of the drug naproxen is a pain reliever and the other enantiomer is a liver toxin. Information about the structure and optical activity of these enantiomers is available online.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>An empirical formula shows the simplest ratio of the elements in a molecule.</p> <p>Elemental microanalysis can be determined from:</p> <ul style="list-style-type: none"> ◆ combustion product masses ◆ percentage product by mass <p>Mass spectrometry can be used to determine the accurate gram formula mass (GFM) and structural features of an organic compound.</p> <p>In mass spectrometry, a small sample of an organic compound is bombarded by high-energy electrons. This removes electrons from the organic molecule generating positively charged molecular ions known as parent ions. These molecular ions then break into smaller positively charged ion fragments. A mass spectrum is obtained showing a plot of the relative abundance of the ions detected against the mass-to-charge (m/z) ratio.</p> <p>The mass-to-charge ratio of the parent ion can be used to determine the GFM of the molecular ion, and so a molecular formula can be determined using the empirical formula.</p> <p>The fragmentation data can be interpreted to gain structural information.</p>	<p>ChemCollective, produced by the National Science Foundations, has a video tutorial: determining the empirical formula from an elemental analysis that shows a worked example of an empirical formula calculation from combustion product masses.</p> <p>Khan Academy has two videos, empirical formula from mass composition and another mass composition problem, that show worked examples of empirical formulae calculations from percentage product by mass.</p> <p>Chemguide has useful information about mass spectrometry including an outline of the process of obtaining a mass spectrum of a sample, fragmentation patterns and molecular ions.</p> <p>A PowerPoint presentation introducing mass spectroscopy is a resource produced by Education Scotland and is available on the Science NQ GLOW portal.</p> <p>Spectroscopy in a suitcase, available through RSC Education, has a Students mass spectrometry exercise providing a forensic basis to analysing mass spectra.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Infrared spectroscopy is used to identify certain functional groups in an organic compound.</p> <p>When infrared radiation is absorbed by organic compounds, bonds within the molecule vibrate (stretch and bend). The wavelengths of infrared radiation that are absorbed depend on the type of atoms that make up the bond and the strength of the bond.</p> <p>In infrared spectroscopy, infrared radiation is passed through a sample of the organic compound and then into a detector that measures the intensity of the transmitted radiation at different wavelengths. The absorbance of infrared radiation is measured in wavenumbers, the reciprocal of wavelength, in units of cm^{-1}.</p> <p>Characteristic absorptions by particular vibrations are given in the data booklet.</p> <p>Proton nuclear magnetic resonance spectroscopy (proton NMR or ^1H NMR) can give information about the different chemical environments of hydrogen atoms (protons or ^1H) in an organic molecule, and about how many hydrogen atoms there are in each of these environments.</p>	<p>A PowerPoint presentation introducing Infrared Spectroscopy is a resource produced by Education Scotland and available on the Science NQ GLOW portal.</p> <p>Chemguide has useful information about infrared spectroscopy including an outline of the process of obtaining an infrared spectrum, fingerprint regions and functional group analysis.</p> <p>RSC Education has a handout, Modern chemical techniques — infrared, that contains some spectra that candidates can analyse.</p> <p>Khan Academy has a number of videos explaining infrared spectroscopy, including practice examples.</p> <p>RSC Education has a handout, Modern chemical techniques — nuclear magnetic resonance spectroscopy, that contains some spectra that candidates can analyse.</p> <p>Khan Academy has a number of videos explaining ^1H NMR spectroscopy, including practice examples.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>^1H nuclei behave like tiny magnets and in a strong magnetic field some align with the field (lower energy), whilst the rest align against it (higher energy). Absorption of radiation in the radio frequency region of the electromagnetic spectrum causes the ^1H nuclei to 'flip' from the lower to the higher energy alignment. As they fall back from the higher to the lower energy alignment the emitted radiation is detected and plotted on a spectrum.</p> <p>In a ^1H NMR spectrum the chemical shift, δ, (peak position) is related to the environment of the ^1H atom and is measured in parts per million (ppm).</p> <p>Chemical shift values for ^1H in different chemical environments are given in the data booklet.</p> <p>The area under the peak is related to the number of ^1H atoms in that environment and is often given by an integration curve on a spectrum. The height of an integration curve is proportional to the number of ^1H atoms in that environment, and so a ratio of ^1H atoms in each environment can be determined.</p> <p>The standard reference substance used in ^1H NMR spectroscopy is tetramethylsilane (TMS), which is assigned a chemical shift value equal to zero.</p> <p>^1H NMR spectra can be obtained using low-resolution or high-resolution NMR.</p>	<p>RSC Education has a handout, Modern chemical techniques — nuclear magnetic resonance spectroscopy, that contains some spectra that candidates can analyse.</p> <p>Khan Academy has a number of videos explaining ^1H NMR spectroscopy, including practice examples.</p> <p>Chemguide has useful information about ^1H NMR spectroscopy, including an outline of the process of obtaining a spectrum of a sample, low- and high-resolution spectroscopy and integration of peaks.</p> <p>Education Scotland has a number of resources available on the Science NQ GLOW portal including:</p> <ul style="list-style-type: none"> ◆ PowerPoint presentations introducing Proton NMR Spectroscopy ◆ learners' workbooks and answers ◆ an organic spectroscopy structural determination workshop that provides an opportunity for candidates to use mass, IR and ^1H NMR spectra to determine the structure of a molecule

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>High-resolution ^1H NMR uses higher radio frequencies than those used in low-resolution ^1H NMR and provides more detailed spectra.</p> <p>In a high-resolution ^1H NMR an interaction with ^1H atoms on neighbouring carbon atoms can result in the splitting of peaks into multiplets. The number of ^1H atoms on neighbouring carbon atoms will determine the number of peaks within a multiplet and can be determined using the $n+1$ rule, where n is the number of ^1H atoms on the neighbouring carbon atom.</p> <p>Low- and high-resolution ^1H NMR spectra can be analysed, and low-resolution ^1H NMR spectra can be sketched for any given compound.</p> <p>(e) Pharmaceutical chemistry</p> <p>Drugs are substances that alter the biochemical processes in the body.</p> <p>Drugs that have beneficial effects are used in medicines.</p> <p>A medicine usually contains the drug plus other ingredients such as fillers to add bulk or sweeteners to improve the taste.</p>	<p>RSC Education, Making medicines video, provides a useful introduction to drug development.</p> <p>Encyclopaedia Britannica has a good introduction to different types of drug action.</p> <p>The articles 'Big picture on drug development' and 'Drug formulation', available through Stem Learning, introduces the process of drug development and formulation.</p> <p>The Association of the British Pharmaceutical Industry has the resource making medicines that gives background information about the drug development process and also has interesting information about the history of medicine through to development of modern medicines.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Drugs generally work by binding to specific protein molecules. These protein molecules can be found on the surface of a cell (receptor) or can be specific enzyme molecules within a cell.</p> <p>Drugs that act on receptors can be classified as agonists or antagonists.</p> <ul style="list-style-type: none"> ◆ An agonist mimics the natural compound and binds to the receptor molecules to produce a response similar to the natural active compound. ◆ An antagonist prevents the natural compound from binding to the receptor, and so blocks the natural response from occurring. <p>Many drugs that act on enzymes are classified as enzyme inhibitors and act by binding to the active site of the enzyme and blocking the reaction normally catalysed there.</p> <p>The overall shape and size of a drug is such that it interacts with a receptor binding site or to the active site of an enzyme. The types of interactions formed can include Van der Waals forces and/or ionic bonds.</p>	<p>The RSC <i>Education in Chemistry</i> article 'Pain relief: from coal tar to paracetamol' discusses the development of paracetamol, its mode of action and the importance of correct doses. <i>The Guardian</i> article, 'Does paracetamol do you more harm than good?', also provides some useful information.</p> <p><i>The Conversation</i> article, 'Explainer: how do drugs work', provides a condensed explanation about agonist and antagonist drugs.</p> <p>Edinformatics has an article that discusses how drugs work and provides examples of enzyme inhibitors, agonist and antagonist drugs, as well as having 3D models of a number of drugs bound to active sites.</p> <p>The British Heart Foundation has short videos with a brief explanation of how aspirin prevents blood clots (enzyme inhibition) forming and how beta blockers work (receptor antagonists).</p> <p>RSC Education has details about aspirin, including a short explanation of its action as an enzyme inhibitor and the history of its development. There are also details of experiments to synthesise, purify, and characterise aspirin that include important practical techniques such as recrystallisation, melting point and thin-layer chromatography analysis.</p> <p>Khan Academy has a video, Beta-lactam antibiotics, which discusses the chemistry of beta-lactam derivatives and how they act as enzyme inhibitors.</p>

Organic chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>The structural fragment of a drug molecule that allows it to form interactions with a receptor binding site or to an enzyme active site normally consists of different functional groups correctly orientated with respect to each other.</p> <p>By comparing the structures of drugs that have similar effects on the body, the structural fragment that is involved in the drug action can be identified.</p>	<p>The article 'Caffeine: The chemistry behind the world's most popular drug' discusses the antagonist nature of caffeine and compares the structure of caffeine to the natural molecule adenosine.</p> <p>RSC Education's Nurofen worksheet provides some information about how Nurofen acts.</p> <p>The website <i>Protein Data Bank</i> has a large number of protein structures to view. For example, the structure of the enzyme neuraminidase, involved in influenza, is shown with its natural substrate as well as with two active drugs.</p> <p>A PowerPoint presentation, Medicinal Chemistry, is a resource produced by Education Scotland and available on the Science NQ GLOW portal. It provides some examples of drugs, their active structural fragments, and their interactions with the active sites.</p> <p>World of Molecules has a selection of 3D drug molecules that can be viewed, as well as an explanation about the mode of action and the history of their development. Another resource available gives the structure of some drug molecules and allows for comparison of the key functional groups necessary for biological activity. Explain it with molecules has information about caffeine.</p>

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(a) Common chemical apparatus</p> <p>Candidates must be familiar with the use(s) of the following types of apparatus:</p> <ul style="list-style-type: none"> ◆ conical flask ◆ digital balance ◆ pipette with safety filler ◆ burette ◆ volumetric (standard) flask ◆ distillation (round-bottomed) flask ◆ condenser ◆ thermometer ◆ Büchner or Hirsch or sintered glass funnel ◆ glassware with ground glass joints ('Quickfit' or similar) ◆ thin-layer chromatography apparatus ◆ colorimeter ◆ melting point ◆ separating funnel 	<p>RSC Education provides details of standard pieces of laboratory equipment in The interactive lab primer — lab apparatus.</p> <p>The interactive lab primer — weighing compounds using a balance, available through RSC Education, contains a video and an online simulation that allows candidates to become familiar with the correct use of chemical balances.</p>

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(b) Skills involved in experimental work Candidates must be able to:</p> <ul style="list-style-type: none"> ◆ tabulate data using appropriate headings and units of measurement ◆ represent data as a scatter graph with suitable scales and labels ◆ sketch a line of best fit (straight or curved) to represent the trend observed in the data ◆ calculate average (mean) values ◆ identify and eliminate rogue points ◆ qualitatively appreciate the relative accuracy of apparatus used to measure the volume of liquids ◆ comment on the reproducibility of results where measurements have been repeated ◆ carry out quantitative stoichiometric calculations ◆ interpret spectral data ◆ appropriately use a positive control, for example a known substance, to validate a technique or procedure 	<p>RSC Education offers a number of activities related to experimental work:</p> <ul style="list-style-type: none"> ◆ a guide to keeping a lab book gives useful advice to candidates about how to keep a good lab book to help with planning their projects ◆ The nature of science: measurement, accuracy and precision supports the teaching of reproducibility, and identifying rogue points and uncertainties <p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland, is available through SSERC and contains useful information about errors and uncertainty calculations (that may be useful for candidates carrying out data analysis as part of their project work).</p> <p>Why do scientists do what scientist do describes and explains how to use a positive control in an experiment. It also gives an introduction to and video on accuracy, precision, errors and statistics.</p> <p>Many of the suggested experiments show the appropriate use of a positive control to validate a technique or procedure. Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, has details of the determination of vitamin C by titration with iodine using a sample of pure ascorbic as a positive control.</p>

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<p>(c) Stoichiometric calculations</p> <p>Stoichiometry is the study of mole relationships involved in chemical reactions.</p> <p>Chemical equations, using formulae and state symbols, can be written and balanced to show the mole ratio(s) of reactants and products, including multi-step reactions.</p> <p>The mass of a mole of any substance, in grams (g), is equal to the gram formula mass (GFM) and can be calculated using relative atomic masses.</p> <p>Calculations can be performed using the relationship between the mass and the number of moles of a substance.</p> <p>For solutions, the mass of solute (grams or g), the number of moles of solute (moles or mol), the volume of solution (litres or l), or the concentration of the solution (moles per litre or mol l⁻¹), can be calculated from data provided.</p> <p>Percentage by mass is the mass of solute made up to 100 cm³ of solution.</p> <p>Percentage by volume is the number of cm³ of solute made up to 100 cm³ of solution.</p> <p>The unit ppm stands for parts per million and refers to 1 mg per kg or 1 mg per litre.</p>	<p>Integrating stoichiometric calculations throughout the course and practical work offers the opportunity to link with real-life examples, leading to deeper understanding.</p> <p>RSC Education, Starters for ten: chapters 1–12, section 2 'Quantitative Chemistry', offers a selection of easily editable short quizzes and activities testing stoichiometric calculations.</p>

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<p>Calculations can be performed using data, including:</p> <ul style="list-style-type: none"> ◆ GFM ◆ masses ◆ number of moles ◆ concentrations and volumes of solutions ◆ volumes of gases ◆ reactant excess ◆ theoretical and percentage yield ◆ empirical formulae <p>Theoretical yields can be calculated and compared with actual yields, leading to determining the percentage yield. The percentage yield is reduced by:</p> <ul style="list-style-type: none"> ◆ mass transfer or mechanical losses ◆ purification of product ◆ side reactions ◆ equilibrium position <p>Candidates must be able to carry out stoichiometric calculations for all of the skills and techniques in the course where appropriate.</p>	<p>Possible experiments to determine number of moles, reactant excess and theoretical and percentage yields include:</p> <ul style="list-style-type: none"> ◆ preparation of aspirin — the RSC offers two aspirin preparations: <ul style="list-style-type: none"> — the Aspirin book contains details of a number of experiments relating to aspirin, including the synthesis, recrystallisation, melting point and TLC analysis of aspirin — The microscale synthesis of aspirin provides instructions for producing small quantities of aspirin that could be used to determine the percentage yield of aspirin produced ◆ Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, has details of the theory of percentage yield and also has instructions for experiments that can be used to practise percentage yield calculations and evaluations: <ul style="list-style-type: none"> — preparation of potassium trioxalatoferrate(III) also provides an opportunity to introduce the practical technique of hot filtration — preparation of benzoic acid by hydrolysis of ethyl benzoate also provides an opportunity to introduce the practical techniques of reflux, recrystallisation, melting point analysis and thin-layer chromatography

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
	<ul style="list-style-type: none"> ◆ the University of Strathclyde provides details of experiments that can be used in practising percentage yield calculations and also involve other practical techniques such as recrystallisation, melting point and thin-layer chromatography: <ul style="list-style-type: none"> — synthesis and analysis of aluminium compounds prepared from Coke cans — electrophilic aromatic substitution organic photosynthesis reactions — preparation of cis- and trans- complexes of cobalt <p>RSC Education, Starters for ten: chapters 1–12, section 2 'Quantitative Chemistry', offers a selection of easily editable short quizzes and activities that includes percentage yield calculations.</p>

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<p>(d) Gravimetric analysis</p> <p>Candidates must be familiar with the technique of gravimetric analysis, including use of:</p> <ul style="list-style-type: none"> ◆ an accurate electronic balance, including the tare function ◆ a weighing boat ◆ weighing by difference ◆ the term ‘weighing accurately approximately’ ◆ heating to constant mass: <ul style="list-style-type: none"> — heating a substance — allowing to cool in a desiccator to prevent absorption of water — weighing — repeating the steps of heating, cooling and weighing until no further changes in mass are observed <p>Gravimetric analysis is used to determine the mass of an element or compound in a substance.</p> <p>The substance is converted into another substance of known chemical composition, which can be readily isolated and purified.</p> <p>The conversion can occur either through precipitation or volatilisation.</p> <p>In precipitation conversion the substance undergoes a precipitation reaction. The precipitate is separated from the filtrate and the filtrate tested to ensure the reaction has gone to completion. The precipitate is washed, dried to constant mass and then weighed.</p>	<p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, contains useful theory about gravimetric analysis as well as details of the following experiments:</p> <ul style="list-style-type: none"> ◆ gravimetric determination of water in hydrated barium chloride by volatilisation ◆ gravimetric determination of Ni using dimethylglyoxime by precipitation <p>Preparation and characterisation of transition metal-oxalate ligand complexes, available through the University of Strathclyde, details the preparation of potassium trioxalatoferrate(III), followed by gravimetric analysis to determine the water of hydration as well as volumetric analysis of oxalate and iron(III) content.</p>

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>In volatilisation conversion the substance is heated and any volatile products (often water) are evaporated. The substance is heated to constant mass and the final mass recorded.</p> <p>(e) Volumetric analysis</p> <p>Candidates must be familiar with use of the technique of volumetric analysis, including:</p> <ul style="list-style-type: none"> ◆ preparing a standard solution ◆ accurate dilution ◆ standardising solutions to determine accurate concentration ◆ titrating to obtain concordancy using burettes, pipettes and volumetric flasks ◆ choosing an appropriate indicator <p>A solution of accurately known concentration is known as a standard solution.</p> <p>A standard solution can be prepared by:</p> <ul style="list-style-type: none"> ◆ weighing a primary standard accurately ◆ dissolving in a small volume of solvent (usually deionised or distilled water) in a beaker ◆ transferring the solution and rinsings into a volumetric flask ◆ making up to the graduation mark with solvent ◆ stoppering and inverting 	<p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, contains useful theory about volumetric analysis as well as details of possible experiments including:</p> <ul style="list-style-type: none"> ◆ preparation of a standard solution of 0.1 mol l⁻¹ oxalic acid ◆ standardisation of a solution of sodium hydroxide using oxalic acid ◆ determination of the ethanoic acid content of vinegar ◆ preparation of a standard solution of 0.1 mol l⁻¹ sodium carbonate ◆ standardisation of a solution of hydrochloric acid using sodium carbonate solution ◆ determination of the purity of marble by back titration ◆ determination of nickel in a nickel(II) salt using EDTA ◆ determination of vitamin C by titration with iodine <p>The water testing resources produced by SSERC contain details of complexometric titrations using EDTA. The Cobalt complexes booklet has details on the preparation and analysis by titration of a cobalt complex and also involves vacuum filtration.</p>

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Standard solutions can also be prepared by accurate dilution by pipetting an appropriate volume of a standard solution into a volumetric flask, making up to the graduation mark with solvent, stoppering and inverting.</p> <p>A primary standard must:</p> <ul style="list-style-type: none"> ◆ be available in a high state of purity ◆ be stable when solid and in solution ◆ be soluble ◆ have a reasonably high GFM <p>Examples of primary standards include:</p> <ul style="list-style-type: none"> ◆ sodium carbonate, Na_2CO_3 ◆ hydrated oxalic acid, $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ◆ potassium hydrogen phthalate, $\text{KH}(\text{C}_8\text{H}_4\text{O}_4)$ ◆ silver nitrate, AgNO_3 ◆ potassium iodate, KIO_3 ◆ potassium dichromate, $\text{K}_2\text{Cr}_2\text{O}_7$ <p>Sodium hydroxide is not a primary standard as it has a relatively low GFM, is unstable as a solid (absorbs moisture) and unstable as a solution. Sodium hydroxide solution must be standardised before being used in volumetric analysis.</p>	<p>RSC Education has a number of resources relating to volumetric analysis:</p> <ul style="list-style-type: none"> ◆ a video with instructions on making a standard solution ◆ the interactive lab primer — standard solution, apparatus guide, and mass and concentration calculators ◆ the interactive lab primer — titration has a video, apparatus guide and information relating to pH curves and indicators as well as instructions for correct use of pipettes and burettes ◆ titration screen experiment provides videos, animations and interactive quizzes relating to titration, including acid-base and redox titrations ◆ a redox titration using wine video and experiment instructions <p>The University of Strathclyde details a number of experiments involving volumetric analysis, including:</p> <ul style="list-style-type: none"> ◆ preparation and characterisation of transition metal-oxalate ligand complexes — preparation of potassium trioxaatoferate(III) followed by gravimetric analysis to determine the water of hydration, as well as volumetric analysis of oxalate and iron(III) content ◆ Build-a-stomach: antacid study — direct and back titration experiment with antacid treatments

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Candidates must be familiar with use of the following types of volumetric analysis:</p> <ul style="list-style-type: none"> ◆ acid-base titrations ◆ redox titrations based on reactions between oxidising and reducing agents ◆ complexometric titrations based on reactions in which complexes are formed — EDTA is an important complexometric reagent and can be used to determine the concentration of metal ions in solution ◆ back titrations used to find the number of moles of a substance by reacting it with an excess volume of a reactant of known concentration. The resulting mixture is then titrated to work out the number of moles of the reactant in excess. From the initial number of moles of that reactant, the number of moles used in the reaction can be determined. The initial number of moles of the substance being analysed can then be calculated. A back titration is useful when trying to work out the quantity of substance in a solid with a low solubility 	

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>(f) Practical skills and techniques</p> <p>Candidates must be familiar with use of the technique of colorimetry, including:</p> <ul style="list-style-type: none"> ◆ preparing a series of standard solutions of appropriate concentration ◆ choosing an appropriate colour or wavelength of filter complementary to the colour of the species being tested ◆ using a blank ◆ preparing a calibration graph 	<p>Candidates can carry out practical skills and techniques as individual activities or combined together during single experiments. Some of the suggestions listed below allow candidates to use a number of practical skills and techniques in a single experiment.</p> <p>The RSC website 'SpectraSchool' provides many resources relating to spectroscopy techniques with animations, videos and interactive spectra. 'Spectroscopy in a suitcase' is an RSC outreach activity giving candidates the opportunity to learn about spectroscopy through hands-on experience. As well as covering the principles of spectroscopic techniques, the activities use real-life contexts to demonstrate the applications of the techniques. You can use this to teach colorimetry. There is also a student resource that introduces spectroscopy as well as providing background information the technique of colorimetry.</p> <p>Challenging medicines resource, available through RSC Education, details colorimetric analysis of 2-hydroxybenzoic acid, aspirin and paracetamol. Also available through RSC Education is challenging plants that details colorimetric analysis of soil samples.</p> <p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, contains useful theory about colorimetric analysis as well as instructions for the determination of manganese in steel experiment.</p>

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<p>Colorimetry uses the relationship between colour intensity of a solution and the concentration of the coloured species present.</p> <p>A colorimeter or a spectrophotometer is used to measure the absorbance of light of a series of standard solutions, and this data is used to plot a calibration graph.</p> <p>The concentration of the solution being tested is determined from its absorbance and by referring to the calibration curve.</p> <p>The concentration of coloured species in the solution being tested must lie in the straight line section of the calibration graph.</p> <p>Candidates must be familiar with use of the technique of distillation. Distillation is used for identification and purification of organic compounds.</p> <p>The boiling point of a compound, determined by distillation, is one of the physical properties that can be used to confirm its identity.</p> <p>Distillation can be used to purify a compound by separating it from less volatile substances in the mixture.</p>	<p>SSERC provides a method for determination of iron and manganese in tea using colorimetry.</p> <p>Colorimetric determination of iron content, available through the University of Strathclyde, provides theory as well as instructions for determination of iron in Irn Bru, iron tablets and spinach.</p> <p>Theory of the organic techniques distillation, heating under reflux, vacuum filtration, solvent extraction, recrystallisation, melting point and mixed melting point determination and thin-layer chromatography are available in Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC.</p> <p>RSC Education provides a number of resources for distillation:</p> <ul style="list-style-type: none"> ◆ interactive lab primer — distillation contains a video, animation and apparatus guide to distillation ◆ extracting limonene from oranges by steam distillation

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Candidates must be familiar with use of the technique of heating under reflux. Heating under reflux allows heat energy to be applied to a chemical reaction mixture over an extended period of time without volatile substances escaping.</p> <p>When carrying out heating under reflux, the reaction mixture is placed in a round-bottomed flask with anti-bumping granules and the flask is fitted with a condenser. The flask is then heated using an appropriate source of heat.</p>	<p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, contains details of two experiments that involve distillation:</p> <ul style="list-style-type: none"> ◆ preparation of ethyl ethanoate ◆ preparation of cyclohexene from cyclohexanol <p>Oils and spices, available through the University of Strathclyde, details the extraction of essential oils by steam distillation followed by solvent extraction.</p> <p>RSC Education provides a number of resources for heating under reflux:</p> <ul style="list-style-type: none"> ◆ interactive lab primer — heating under reflux has a video, animation and apparatus guide ◆ the synthesis of aspirin also provides an opportunity to introduce other important practical techniques including recrystallisation, vacuum filtration, melting point analysis, thin-layer chromatography and % yield calculations. A similar procedure is available in Chemistry: A Guide to Practical Support Materials (2012) ◆ the RSC aspirin screen experiment gives an opportunity to learn about the synthesis of aspirin before beginning practical work. It provides videos, animations and interactive quizzes relating to the synthesis of aspirin, purification by recrystallisation and analysis by TLC and relates these experiments to the functional groups and properties of aspirin

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
	<ul style="list-style-type: none"> ◆ The First Year Undergraduate Chemistry Laboratory Course Manual 2011–2012 provides experimental details for the preparation of ester propyl ethanoate and the preparation of benzoic acid from methyl benzoate, and also provides an opportunity to include other important practical techniques (vacuum filtration, recrystallisation, melting point analysis and % yield calculations) <p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, contains details of an experiment to hydrolyse ethyl benzoate, by heating under reflux, to prepare benzoic acid. It also provides an opportunity to include other important practical techniques (recrystallisation, vacuum filtration, melting point analysis, thin-layer chromatography and % yield calculations). This guide also contains experimental details for preparation of ethyl ethanoate.</p> <p>The University of Strathclyde details a number of experiments involving heating under reflux:</p> <ul style="list-style-type: none"> ◆ preparation and characterisation of methyl esters details experimental procedures to make two solid methyl esters, and involves the practical techniques of solvent extraction and melting point determination as well as providing some spectral data ◆ using an inorganic complex to catalyse an organic molecular reaction provides an opportunity to carry out vacuum filtration, solvent extraction and thin-layer chromatography

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Candidates must be familiar with use of the technique of vacuum filtration. Vacuum filtration involves carrying out a filtration under reduced pressure and provides a faster means of separating a precipitate from a filtrate. A Büchner, Hirsch or sintered glass funnel can be used during vacuum filtration.</p>	<p>RSC Education's interactive lab primer — vacuum filtration contains a video and an apparatus guide.</p> <p>There are many experiments that could include vacuum filtration. A selection provided in Chemistry: A Guide to Practical Support Materials (2012) includes:</p> <ul style="list-style-type: none"> ◆ preparation of potassium trioxalatoferrate(III) ◆ preparation of aspirin ◆ preparation of benzoic acid by hydrolysis of ethyl benzoate <p>The University of Strathclyde also details the following experiments that use vacuum filtration:</p> <ul style="list-style-type: none"> ◆ cis- and trans-complexes of cobalt ◆ methyl orange synthesis ◆ electrophilic aromatic substitution reactions provides an opportunity to introduce the practical techniques of recrystallisation, melting point analysis and thin-layer chromatography

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Candidates must be familiar with use of the technique of recrystallisation to purify an impure solid involving:</p> <ul style="list-style-type: none"> ◆ dissolving an impure solid gently in a minimum volume of a hot solvent ◆ hot filtration of the resulting mixture to remove any insoluble impurities ◆ cooling the filtrate slowly to allow crystals of the pure compound to form, leaving soluble impurities dissolved in the solvent ◆ filtering, washing and drying the pure crystals <p>The solvent for recrystallisation is chosen so that the compound being purified is completely soluble at high temperatures and only sparingly soluble at lower temperatures.</p> <p>Candidates must be familiar with use of the technique of solvent extraction. Solvent extraction involves isolating a solute from a liquid mixture or solution by extraction using an immiscible solvent in which the solute is soluble.</p>	<p>In addition to the previous suggested experiments that could include recrystallisation, RSC Education also provides:</p> <ul style="list-style-type: none"> ◆ interactive lab primer – recrystallisation, including a video, animation and an apparatus guide as well as details about hot filtration ◆ preparation of paracetamol <p>There are many experiments that could include recrystallisation. A selection provided in Chemistry: A Guide to Practical Support Materials (2012) includes:</p> <ul style="list-style-type: none"> ◆ preparation of aspirin ◆ preparation of benzoic acid by hydrolysis of ethyl benzoate <p>RSC Education provides a number of resources for solvent extraction including:</p> <ul style="list-style-type: none"> ◆ interactive lab primer — solvent extraction, which has a video, animation and apparatus guide

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>When carrying out a solvent extraction, the two immiscible solvents form two layers in the separating funnel. The solute dissolves in both solvents and an equilibrium establishes between the two layers. The ratio of solute dissolved in each layer is determined by the equilibrium constant, K. The lower layer is run off into a container and the upper layer is poured into a second container. This process is repeated to maximise the quantity of solute extracted.</p> <p>The quantity of solute extracted is greater if a number of extractions using smaller volumes of solvent are carried out rather than a single extraction using a large volume of solvent.</p> <p>The solvent used should be:</p> <ul style="list-style-type: none"> ◆ immiscible with the liquid mixture or solution (usually water) ◆ one in which the solute is more soluble in than the liquid mixture or solution (usually water) ◆ volatile to allow the solute to be obtained by evaporation of the solvent ◆ unreactive with the solute 	<ul style="list-style-type: none"> ◆ extracting iodine from seaweed by solvent extraction (only one extraction is carried out in this procedure) ◆ preparation of cyclohexene from cyclohexanol with purification by distillation and solvent extraction <p>Chemistry: A Guide to Practical Support Materials (2012), produced by Education Scotland and available through SSERC, details experiments that provide an opportunity to introduce the practical techniques of distillation and solvent extraction (only one extraction is carried out in both of these experiments):</p> <ul style="list-style-type: none"> ◆ preparation of cyclohexene from cyclohexane using concentrated phosphoric acid ◆ preparation of ethyl ethanoate <p>The University of Strathclyde also details the following experiments that use solvent extraction:</p> <ul style="list-style-type: none"> ◆ extraction of caffeine from tea provides an opportunity to perform melting point analysis and percentage yield calculations ◆ plant pigments and pH indicators provides an opportunity to carry out thin-layer chromatography ◆ selective reductions with sodium borohydride provides an opportunity to monitor reactions using thin-layer chromatography

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Candidates must be familiar with use of the techniques of melting point and mixed melting point determination. The melting point of a substance is the temperature range over which the solid first starts to melt, to when all of the solid has melted.</p> <p>The identity of a pure compound can be confirmed by melting point analysis and a comparison of the experimentally determined melting point with a literature or known melting point value.</p> <p>Determination of the melting point of a compound can give an indication of the purity of a compound. The presence of impurities in the compound lowers the melting point and broadens its melting temperature range due to the disruption in intermolecular bonding in the crystal lattice.</p> <p>Determination of a mixed melting point involves mixing a small quantity of the product with some of the pure compound and determining the melting point. The melting point value and the range of the melting temperature can be used to determine if the product and the pure compound are the same substance.</p> <p>Candidates must be familiar with use of the technique of thin-layer chromatography. Chromatography is a technique used to separate the components present within a mixture. Chromatography separates substances by making use of differences in their polarity or molecular size.</p>	<p>In addition to the previous suggested experiments that could include melting point determination, RSC Education's interactive lab primer has a video and apparatus guide to melting point determination.</p> <p>As well as learning about the practical technique of melting point analysis, candidates could explain the effect impurities have on the melting point of a solid in terms of the strengths of intermolecular forces of attraction. The University of Rhode Island explains the effect impurities have on melting point, in relation to strength of intermolecular forces.</p>

Researching chemistry Mandatory knowledge	Suggested learning activities and resources, and/or further guidance on mandatory knowledge
<p>Thin-layer chromatography (TLC) uses a fine film of silica or aluminium oxide spread over glass, aluminium foil or plastic. A small sample of the mixture being tested is spotted onto the base (pencil) line of the chromatogram. A solvent dissolves the compounds in the spot and carries the compounds up the chromatogram. How far the compounds are carried depends on how soluble the compounds are in the chosen solvent and how well they adhere to the plate. A developing agent or ultraviolet light is normally required to visualise the spots on the chromatogram.</p> <p>R_f values can be calculated:</p> $R_f = \frac{\text{distance travelled by the sample}}{\text{distance travelled by the solvent}}$ <p>Under the same conditions (temperature, solvent, and saturation levels) a compound always has the same R_f value (within experimental error).</p> <p>The identity of a compound can be confirmed by:</p> <ul style="list-style-type: none"> ◆ comparing the experimentally determined R_f values with a literature or known value determined under the same conditions ◆ making a direct comparison on a TLC plate between the compound being tested and the pure substance — a co-spot could be used <p>TLC is used to assess the purity of substances. A pure substance, when spotted and developed on a TLC plate, should appear as a single spot (some impurities may not be visible by TLC analysis). The presence of more than one spot shows that impurities are present.</p>	<p>In addition to the previous suggested experiments that could include TLC, RSC Education's interactive lab primer has a video, animation and apparatus guide to thin layer chromatography.</p> <p>The University of Strathclyde details electrophilic aromatic substitution reactions that include instructions for carrying out TLC, as well as providing IR spectra for analysis.</p>

Preparing for course assessment

The course is assessed by a question paper and a project. To help candidates prepare for the course assessment, you should give them opportunities to practise activities similar to those expected in the course assessment, for example:

- ◆ practising question paper techniques and revising for the question paper. To support this learning, you may find it helpful to refer to: *Advanced Higher Chemistry Specimen Question Paper*
- ◆ preparing for the project: selecting topics; gathering and researching information/data; evaluating and analysing findings; developing and justifying conclusions; and presenting the information/data (as appropriate). You may find it helpful to refer to the *Advanced Higher Chemistry coursework assessment task*

Developing skills for learning, skills for life and skills for work

You should identify opportunities throughout the course for candidates to develop skills for learning, skills for life and skills for work.

Candidates should be aware of the skills they are developing and you can provide advice on opportunities to practise and improve them.

Qualifications Scotland does not formally assess skills for learning, skills for life and skills for work.

There may also be opportunities to develop additional skills depending on the approach centres use to deliver the course. This is for individual teachers and lecturers to manage.

Some examples of potential opportunities to practise or improve these skills are provided below.

Literacy

Writing means the ability to create texts which communicate ideas, opinions and information, to meet a purpose and within a context. In this context, 'texts' are defined as word-based materials (sometimes with supporting images) which are written, printed, Braille or displayed on screen. These are technically accurate for the purpose, audience and context.

1.1 Reading

Candidates understand and interpret a variety of scientific texts.

1.2 Writing

Candidates use skills to effectively communicate key areas of chemistry, make informed decisions and explain, clearly, chemistry issues in various media forms. Candidates have the opportunity to communicate applied knowledge and understanding throughout the course.

There are opportunities to develop the literacy skills of listening and reading when gathering and processing information in chemistry.

Numeracy

This is the ability to use numbers in order to solve problems by counting, doing calculations, measuring, and understanding graphs and charts. This is also the ability to understand the results.

Candidates extract, process and interpret information presented in numerous formats including tabular and graphical. Practical work provides opportunities to develop time management and measurement skills.

2.1 Number processes

Number processes means solving problems through: carrying out calculations, when dealing with data and results from experiments/investigations and class work; making informed decisions based on the results of these calculations, and understanding these results.

2.2 Money, time and measurement

The accuracy of measurements is important when handling data in a variety of chemistry contexts, including practical and investigative. Candidates should consider uncertainties.

2.3 Information handling

Candidates experience information handling opportunities when dealing with data in tables, charts and other graphical displays, to draw conclusions with justifications throughout the course. This involves interpreting the data, and considering its reliability in making reasoned deductions and informed decisions with justifications.

Thinking skills

This is the ability to develop the cognitive skills of remembering and identifying, understanding and applying.

The course allows candidates to develop skills of applying, analysing and evaluating. Candidates can analyse and evaluate practical work and data by reviewing the process, identifying issues and forming valid conclusions. They can demonstrate understanding and application of concepts and explain and interpret information and data.

5.3 Applying

Candidates plan experiments and investigations throughout the course and use existing information to solve problems in different contexts.

5.4 Analysing and evaluating

During practical work, candidates review and evaluate experimental procedure and identify improvements. Candidates use their judgement when drawing conclusions from experiments.

Analysis is the ability to solve problems in chemistry and make decisions that are based on available information.

It may involve reviewing and evaluating relevant information and/or prior knowledge to provide an explanation.

It may build on selecting and/or processing information, so is a higher skill.

5.5 Creating

This is the ability to design something innovative or to further develop an existing thing by adding new dimensions or approaches. Candidates can demonstrate creativity, in particular, when planning and designing experiments/investigations. They have the opportunity to be innovative in their approach, and to make, write, say or do something new.

Candidates also have opportunities to develop the skills of working with others, creating, and citizenship.

Working with others

Learning activities provide many opportunities, in all areas of the course, for candidates to work with others. Practical activities and investigations, in particular, offer opportunities for group work, which is an important aspect of chemistry.

Citizenship

Candidates develop citizenship skills, when considering the applications of chemistry on our lives, as well as the implications for the environment and society.

Administrative information

Published: June 2026 (version 4.0)

History of changes

Version	Description of change	Date
2.0	<p>Amendments made in 'Skills knowledge and understanding' section:</p> <ul style="list-style-type: none">◆ 'Inorganic Chemistry' section, sub-section '(c) transitional metals': information about heterogeneous catalysts, heterogeneous catalysis and homogeneous catalysis clarified.◆ 'Physical Chemistry' section, sub-section '(b) reaction feasibility': error in entropy equation corrected.◆ 'Organic Chemistry and instrumental analysis' section, sub-section '(b) synthesis': reference to Markonikov's rule clarified. <p>Course support notes added as appendix.</p>	September 2019
3.0	<p>Error in terminology corrected on p 9 and p 59 — '...in a molecule or neutral ion' changed to '...in a neutral compound'.</p>	October 2020
3.1	<p>Fixed software formatting error that occurred in version 3.0 on p 12 and p 15.</p>	April 2021
4.0	<p>We've updated the 'Course overview' and 'Course assessment' sections to reflect changes to the question paper. We've reduced the question paper from 110 marks to 100 marks, and we'll no longer apply scaling. (Section 1 still has 25 marks, and section 2 now has 75 marks instead of 85.) We've also reduced the duration of the question paper from 3 hours to 2 hours and 30 minutes.</p> <p>The project still has 25 marks, and we'll scale this to 34 marks instead of 40.</p> <p>We've updated the table in the 'Course assessment structure: project' section so that the information in the 'Expected response' column matches the wording in the project assessment task.</p>	June 2026

Version	Description of change	Date
	<p>In the course support notes:</p> <ul style="list-style-type: none"> ◆ We've updated the hyperlinks in the 'Suggested learning activities and resources, and/or further guidance on mandatory knowledge' column in the 'Approaches to learning and teaching' table in the course support notes. ◆ We've corrected typographical errors on pages 95 and 108: 'sulphuric' to 'sulfuric' and 'neurominidase' to 'neuraminidase'. ◆ We've updated the 'Preparing for course assessment' section to remove reference to an outdated document. <p>We've changed references to 'SQA' to 'Qualifications Scotland' throughout. We've updated the title to include the SCQF level.</p> <p>We've updated the 'Equality and inclusion' and 'Further information' sections and added a new 'Learning for Sustainability' section.</p> <p>We've made changes to the format and layout of the document to improve accessibility.</p> <p>What you need to do differently</p> <p>There is no change to teaching and learning. You can update your assessments, for example prelims, to reflect the changes to the question paper.</p>	

Note: please check Qualifications Scotland's website to ensure you are using the most up-to-date version of this document.

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